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Application of several activity coefficient models to water-organic-electrolyte aerosols of atmospheric interest

T. Raatikainen and A. Laaksonen

Department of Applied Physics, University of Kuopio, P.O. Box 1627, 70211 Kuopio, Finland

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Correspondence to: T. Raatikainen (tomi.raatikainen@uku.fi)

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Abstract

In this work, existing and modified activity coefficient models are examined in order to assess their capabilities to describe the properties of aqueous solution droplets relevant in the atmosphere. Five different water-organic-electrolyte activity coefficient models were first selected from the literature. Only one of these models included organics and electrolytes which are common in atmospheric aerosol particles. In the other models, organic species were solvents such as alcohols, and important atmospheric ions like NH_4^+ could be missing. The predictions of these models were compared to experimental activity and solubility data in aqueous single electrolyte solutions with 31 different electrolytes.

Based on the deviations from experimental data and on the capabilities of the models, four predictive models were selected for fitting of new parameters for binary and ternary solutions of common atmospheric electrolytes and organics. New electrolytes (H^+, NH_4^+, Na^+, Cl^-, NO_3^- and SO_4^{2-}) and organics (dicarboxylic and some hydroxy acids) were added and some modifications were made to the models if it was found useful. All new and most of the existing parameters were fitted to experimental single electrolyte data as well as data for aqueous organics and aqueous organic-electrolyte solutions. Unfortunately, there are very few data available for organic activities in binary solutions and for organic and electrolyte activities in aqueous organic-electrolyte solutions. This reduces model capabilities in predicting solubilities.

After the parameters were fitted, deviations from measurement data were calculated for all fitted models, and for different data types. These deviations and the calculated property values were compared with those from other non-electrolyte and organic-electrolyte models found in the literature. Finally, hygroscopic growth factors were calculated for four 100 nm organic-electrolyte particles and these predictions were compared to experimental data and to predictions from other models.

All of the newly fitted models show good agreement with experimental water activity data in binary and ternary solutions. One of the models is for activities of non-

electrolytes only, but the other three models show quite small deviations from measured electrolyte activities. Because there were not enough experimental data for organic and electrolyte activities, some models show bigger deviation for mutual deliquescence relative humidities of organic-electrolyte particles, but calculated growth factors for liquid droplets are quite close to the experimental data. Even in cases with somewhat bigger deviations, the results can be considered satisfactory, because they were calculated based mainly on the predictive properties of the models.

1. Introduction

Atmospheric aerosols have a cooling effect on the climate, but the magnitude of this effect is uncertain. Aerosol composition and size distribution are the most important properties that influence the magnitude of the cooling. Hygroscopic growth of aerosols and partitioning of volatile compounds between gas and particulate phases can be calculated with thermodynamic models in which non-ideal behaviour of liquid phase compounds are described with modelled activity coefficients. Most of the activity coefficient models used in aerosol modelling are suitable only for electrolyte solutions. However, atmospheric aerosols include both organic and inorganic components (Novakov and Penner, 1993; Saxena and Hildemann, 1996; Chow et al., 1994).

Water and electrolyte activities are calculated in most of the aerosol models, but the organic fraction is usually treated as an ideal solute or insoluble solid phase, or organics can be totally ignored. These are valid approximations in solutions where the organic fraction is small. However, it has been estimated that up to 50% of aerosol dry mass is organic. Even if variations for the organic mass fractions are large, ignoring the organics can cause large deviations on calculated results. An activity coefficient model which is applicable for aqueous mixtures of organics and electrolytes is needed. The problem is that there are only a few activity coefficient models for organic-electrolyte mixtures (e.g. Ming and Russell, 2002). The two main reasons for this are lack of a reliable activity coefficient model for any kind of organic-electrolyte mixtures, and lack

of experimental data for the atmospheric mixtures.

Because electrolytes and organics have different interactions in aqueous solutions, a large fraction of all activity coefficient models are applicable either for non-electrolyte solutions (e.g. Fredenslund et al., 1975; Abrams and Prausnitz, 1975) or for aqueous electrolyte solutions only (e.g. Clegg et al., 1998a; Pitzer, 1991). There are some methods for calculating activities in mixed solutions from binary solution data (e.g. Clegg et al., 2001; Topping et al., 2005). However, the results can be better if organic-electrolyte interactions are taken into account directly.

Most of the few organic-electrolyte models are designed for industrial purposes (e.g. Li et al., 1994), and the compounds in these models are usually different from the major species found in atmospheric aerosols. However, if a suitable model is found, new parameters can be fitted for the atmospheric compounds provided that experimental data is available.

Some activity coefficient models are designed for one purpose only (e.g. vapour-liquid equilibrium models). However, most organics and electrolytes form a solid phase at low relative humidities, and some of these compounds can also evaporate from aqueous droplets. For these reasons, the activity coefficient model should be applicable for water, organic and electrolyte activities in dilute and in saturated solutions.

Usually the inorganic fraction in atmospheric particles is composed of a few different ions (e.g. sulphates, nitrate, hydrogen and ammonium) whose chemical properties are well known. Several studies show that dicarboxylic acids are among the most common atmospheric organics, however, a number of other important species exist as well (Hemming and Seinfeld, 2001). Most of these organics have complex structures, and usually they are less studied than the inorganic compounds. Although there are some data sets available for the atmospheric organics, there are hardly any data for the mixtures of organics and electrolytes. However, there are predictive activity coefficient models that are based on the group contribution method, in which several similar compounds can be described with the same interaction parameters. Experimental data in parameter fitting can include only some compounds from one group, but the fitted

parameters are applicable for the whole group. Then activities of the other compounds can be predicted. UNIFAC (Fredenslund et al., 1975) is one very popular predictive non-electrolyte activity coefficient model.

The first purpose of this study was to compare existing activity coefficient models that might be suitable for modelling of the hygroscopic properties of the atmospheric organic-electrolyte particles. The second purpose was to select some of these models to be extended for aqueous organic-electrolyte solutions of atmospheric interest by fitting new parameters. Because the experimental data needed for the fitting is very limited, we selected models that are predictive. The electrolytes in these models are composed of the ions H^+, NH_4^+, Na^+, Cl^-, NO_3^- and SO_4^{2-}. The organics selected for the parameter fitting are dicarboxylic acids (from oxalic to adipic acid) and acids including hydroxyl groups (citric, tartaric and malic acids). Because the fitted models are predictive, activities can be calculated for other similar multifunctional organic compounds.

2. Thermodynamics

In non-ideal solutions, activities (a) represent the effective concentrations which should be used instead of the real concentrations (c). The relation between the effective and the real concentration is given by the activity coefficient (γ): $a=\gamma c$. Because there are different concentration scales, activities and activity coefficient depend on the concentration scale. A common unit for expressing electrolyte concentration is molality m (mol/kg) i.e. number of electrolyte moles per one kilogram of solvent, where solvent can be defined as a single compound (e.g. water) or a solvent mixture (e.g. water-ethanol solution). The mole fraction x is commonly used for non-electrolyte solutions.

It should be noted that organic compounds can, depending on the species, be considered either solvents or solutes. Some organics can, for example, dissolve electrolytes, whereas others are sparingly soluble in water. From the modelling viewpoint this is an important difference, as the standard state (at which thermodynamic poten-

tials are compared) and the reference state (ideal solution limit) are usually different for solutes and solvents. In the following calculations, organics are considered as solvents.

The standard state is an ideal solution (γ=1) with unit concentration at 298.15 K. For solvents this is taken to be pure ideal solvent and for solutes ideal 1 molal aqueous solution. Pure water can be considered as an ideal solution, but usually 1 molal electrolyte solutions are not ideal, so this state is hypothetical. The selected organics are solids at room temperature, so also their standard states are hypothetical. The reference state for solvents is pure solvent and the reference state for electrolytes is infinitely dilute aqueous solution.

Water in a droplet is in equilibrium with ambient water vapour, when the vapour pressure at droplet surface equals the atmospheric partial pressure. For large droplets (larger than 100 nm in diameter), a good approximation is that the vapour pressure at droplet surface equals the saturation vapour pressure multiplied by water activity. By combining these considerations with the definition of relative humidity (RH), we have $a_w = RH/100\%$. However, according to the Kelvin equation, the equilibrium vapour pressure over a curved surface is higher than that over a flat surface. The equilibrium relative humidity is now

$$RH/100\% = a_w \exp\left(\frac{4\sigma M_w}{RT\rho D_p}\right) \tag{1}$$

where σ (N/m) is surface tension, M_w (kg/mol) is molecular weight for water, R=8.31451 J/(mol K) is gas constant, T (K) is temperature, ρ (kg/m^3) is solution density and D_p (m) is droplet diameter. Similarly, the vapour-liquid equilibrium (VLE) could be calculated for all volatile species including ammonia, some acids (e.g. HCl and HNO_3) and some of the organics.

Most organic solvents are liquids at room temperature (e.g. ethanol), but in this study all pure organics are solids at room temperature. The melting temperatures (Table 1) can be much higher than the boiling point of water. Most electrolytes are also solids at room temperature with the exception of strong acids. For example, sulphuric acid does

not have a solid phase at room temperature. Solid-liquid equilibrium (SLE) is strongly temperature dependent, and also here, solubilities are calculated as a function of temperature. The experimental solubility data used in this work falls in the temperature range 273.15–373.15 K.

Salt solubilities can be calculated using solubility products K_{sp}, i.e. the product of cation (c) and anion (a) activities in the saturated solution.

$$K_{sp} = (a_c^{sat})^{\nu+}(a_a^{sat})^{\nu-} \tag{2}$$

where $\nu+$ and $\nu-$ are the stoichiometric numbers of cations and anions in the salt. The solubility product is a function of temperature, but usually tabulated K_{sp} values are given at 298.15 K. In addition to tabulated values, K_{sp} can be calculated from the change of the Gibbs free energy using the free energies of formation ΔG_f (J/mol) for solid ($\Delta G_{f,(s)}$) and dissolved salts ($\Delta G_{f,(aq)}=(\nu+)\Delta G_{f,c} + (\nu-)\Delta G_{f,a}$):

$$K_{sp}(T) = \exp\left(\frac{\Delta G_{f,(s)} - \Delta G_{f,(aq)}}{RT}\right) \tag{3}$$

These free energies of formation are tabulated at standard temperature (T_0), usually 298.15 K (ΔG_f^0). The temperature dependence of standard state chemical potentials can be calculated with an equation given by Ansari and Pandis (1999). With these standard state chemical potentials we can calculate the free energy change as a function of temperature. By combining Eq. (3) and the equation for the temperature dependence of the free energy change, we obtain an equation for the temperature dependence of solubility products

$$K_{sp}(T) = K_{sp}(T_0)\exp\left[\left(\left(H^0_{f,(s)} - H^0_{f,(aq)}\right)\left(\frac{1}{T} - \frac{1}{T_0}\right) + \left(C_{(s)} - C_{(aq)}\right)\left(\ln\left(\frac{T_0}{T}\right) - \frac{T_0}{T} + 1\right)\right)\frac{1}{R}\right] \tag{4}$$

where $H^0_{f,(s)}$ and $H^0_{f,(aq)}$ (J/mol) are the enthalpies of formation of solid salt and aqueous ions at 298.15 K respectively, and $C_{(s)}$ and $C_{(aq)}$ (J/(mol K)) are the heat capacities for solid salt and aqueous ions respectively. Solubility products at 298.15 K ($K_{sp}(T_0)$) for the six salts modelled in the fitted models are from Clegg et al. (1998b) and the other thermodynamic constants are from Weast (1987) and Ansari and Pandis (1999). If the solubility products at 298.15 K were not available (comparison of existing models), these were calculated from the standard state Gibbs free energies of formation with Eq. (3).

There are no solubility products available for organics, but saturation activities (a_o^{sat}) can be approximated using the enthalpy of melting ΔH_m (J/mol) and melting temperature T_m (K)

$$\ln a_o^{sat} = \frac{\Delta H_m}{R}\left(\frac{1}{T_m} - \frac{1}{T}\right) \tag{5}$$

Because melting enthalpy is a function of temperature, this equation is more accurate if the temperature T is close to the melting temperature. However, most of the melting temperatures are higher than 400 K. There were no reliable melting enthalpies available for oxalic and citric acids so these values were calculated based on thermodynamic data from the NIST Chemistry webBook (http://webbook.nist.gov) and Knovel database (http://www.knovel.com/knovel2/default.jsp). Because these values may be inaccurate, the solubilities of oxalic and citric acids were not used in the fitting process. However, the solubilities were included in the comparisons of the fitted models. Enthalpies of phase transitions between different crystalline forms were included in the melting enthalpies if these values were available. Melting enthalpies and melting temperatures for the modelled organics are given in the Table 1.

2.1. Growth factors for organic-salt particles

Hygroscopic growth factors (GF, aqueous droplet diameter divided by dry particle diameter) can be calculated for originally dry particles (known size and organic-salt mass

ratio) as a function of ambient relative humidity by solving the liquid phase equilibrium concentrations. Pure component densities and molecular weights, which are needed in solving mole numbers from dry particle volume and organic-salt mass ratio, are given in the Table 1 for organics. The corresponding electrolyte parameters can be found from Weast (1987).

Droplet volume is proportional to the cube of the diameter, and thus the growth factor can be calculated from dry and wet particle volumes. We can calculate the dry volume from the given dry diameter, but direct calculation of the liquid phase volume is difficult without solution density data. For simplicity, we approximated the liquid droplet volume as sum of the dry volume and the volume of condensed water.

The equilibrium liquid phase concentrations can be calculated by solving the water concentration so that the Kelvin equation (Eq. 1) is satisfied. At the same time organic and salt concentrations in the liquid phase must be less than or equal to their saturation concentrations. If there is a solid phase, the corresponding liquid phase compound must be saturated. The solid-liquid equilibrium is found by transferring organic compound and salt between the liquid and the solid phases while maintaining the total concentration the same as in the original dry particle.

If a solid phase exists, there exist a single relative humidity at which both the organic compound and the salt are saturated. This relative humidity, called mutual deliquescence relative humidity (*MDRH*), is the minimum relative humidity where the liquid phase exists. Only supersaturated solutions can exist below the *MDRH*. Also, there can be a relative humidity at which the solid phase dissolves completely. After this *RH* the solid phase is not possible, and the equilibrium water concentration is calculated by solving Eq. (1) only.

Solution density and surface tension are needed in Eq. (1), but these are rarely available for multicomponent mixtures. For simplicity, we decided to use the surface tension (0.072 N/m) and the density of pure water (997.1 kg/m^3). Growth factors were calculated for particles with dry diameter of 100 nm, so the Kelvin term has only a small effect to the equilibrium vapour pressure.

3. Comparison of published activity coefficient models

Five different activity coefficient models were selected from the literature. All models are suitable for water-organic-electrolyte mixtures. However, in four of these models the organic species are alcohols and other organic solvents which are not common in atmospheric aerosols. Moreover, some of the important atmospheric ions (e.g. H^+ and NH_4^+) are missing from most of the models. Only the fifth model is fitted for organic-electrolyte mixtures relevant in the atmosphere, however, parameters are available for a few ternary mixtures only.

Some of the selected models have a commonly used name, but not all. Here, the models are called LIQUAC (Li et al., 1994), LIFAC (Yan et al., 1999), Extended UNIQUAC (Thomsen et al., 1996), Modified UNIFAC (Achard et al., 1994) and Ming and Russell model (Ming and Russell, 2002), which is the only model fitted for atmospheric mixtures.

These models are combinations of different parts, usually an electrolyte part, a non-electrolyte part and an additional mixture part. LIQUAC has three parts, where the electrolyte part is based on the Debye-Hückel theory (Fowler and Guggenheim, 1949), the non-electrolyte part is UNIQUAC (Abrams and Prausnitz, 1975) and the third part is virial equation for solvent-ion interactions. LIFAC is similar to LIQUAC except that all species are described with functional groups. For this reason the non-electrolyte part is UNIFAC (Fredenslund et al., 1975) instead of UNIQUAC. Extended UNIQUAC is a combination of UNIQUAC as non-electrolyte part and an electrolyte part based on Debye-Hückel theory. Modified UNIFAC has Larsen's UNIFAC (Larsen et al., 1987) as non-electrolyte part and a Debye-Hückel theory based electrolyte part. Ming and Russell model is a combination of UNIFAC as non-electrolyte part and the Clegg et al. (1992) electrolyte model. In fact, Ming and Russell model reduces to the Clegg model in aqueous electrolyte solutions.

All necessary parameters for LIQUAC, LIFAC and Modified UNIFAC have been published in the original articles. Parameters for Extended UNIQUAC are published in

Thomsen et al. (1996); Iliuta et al. (2000); Thomsen and Rasmussen (1999). Parameters for the non-electrolyte part of the Ming and Russell model are given in the original article. The electrolyte part, i.e. the Clegg model, has temperature dependent parameters for four ions corresponding to four single electrolyte solutions (Clegg et al., 1998a). However, because parameters for most of the electrolytes are available only at constant temperature (298.15 K), we decided to use these constant temperature parameters for all electrolytes. The model parameters are from Clegg et al. (1992, 1998b).

Some of the models have limitations, such as vapour-liquid equilibrium only, room temperature only, maximum electrolyte molality, etc., but in the model comparison these limitations are ignored. The only limitation we apply is that maximum molality is restricted to 30 mol/kg for all electrolytes. In fact, most of the published models are for vapour-liquid equilibrium, but equations are also given for electrolyte activities. Some of these models include values that are functions of temperature (e.g. water density, dielectric constant), but here these are considered constant. This approximation has very small effect to the calculated activity coefficients.

Most electrolytes dissociate completely, but one exception is sulphuric acid (H_2SO_4), for which the second dissociation ($HSO_4^- \rightleftharpoons H^+ + SO_4^{2-}$) is not complete. The equilibrium could be calculated, but this is done only in the Ming and Russell (Clegg) model. In the other models, sulphuric acid is taken to dissociate completely into two H^+ and one SO_4^{2-} ions.

3.1. Experimental single electrolyte data

The only common solvent for the five models is water and for this reason, the models were tested only for aqueous single electrolyte solutions. Because there are only few electrolyte solutions for which all models are applicable, a large number of electrolytes was selected for the model comparison. The 31 electrolytes are H_2SO_4, HNO_3, HCl, $(NH_4)_2SO_4$, NH_4NO_3, NH_4Cl, Na_2SO_4, $NaNO_3$, NaCl, Li_2SO_4, K_2SO_4, NaBr, NaOH, NaI, LiCl, KCl, $CaCl_2$, $Ca(NO_3)_2$, KBr, LiBr, $CaBr_2$, $ZnSO_4$, KNO_3, BaI_2, $BaBr_2$, KI, KF, $LiNO_3$, $ZnCl_2$, $CuCl_2$ and LiI.

Experimental single electrolyte data is categorized into five different data types. Each data type includes at least six but not more than 31 experimental data sets for single electrolytes. Data types and data sources for each electrolyte are presented in Table 2.

Mean activity coefficient and water activity data is given at constant temperature, which is usually 298.15 K, but there are some data sets measured at 293.15 K and 303.15 K. Electrolyte molalities depend on the data set, but usually the molalities range from zero to saturation molality. Solubility and water saturation activity data is given as a function of temperature. Temperatures, as well as solubilities, depend on the data set, but the minimum temperature is 273.15 K and the maximum temperature is 373.15 K.

3.2. Deviations from experimental data

Model deviations were calculated for each of the five data types. Model deviation for one data type is calculated as average of the deviations of its data sets. The deviation for one data set is the average absolute difference between model predictions and experimental data. Because these models are suitable for different mixtures, the number of included data sets is given for each model and data type. Results of the model comparisons are given in Table 3.

When deviations were calculated for the Modified UNIFAC, it was found that there are three mean activity coefficient data sets for which the average absolute deviation is more than 100. It was considered that the Modified UNIFAC is not suitable for these mixtures so these three average absolute deviations were not included in the model deviation.

Modified UNIFAC uses hydrated mole fractions, meaning that some fraction of the liquid water is assumed to be incorporated in hydrates. For this reason, it is possible that all water will be in hydrates at high electrolyte molalities resulting in zero water mole fraction whereby activity coefficients can not be calculated. This characteristic makes it difficult to use the model for very soluble electrolytes (e.g. sulphuric acid). However, this model includes a large number of different electrolytes, and most of these are not

extremely soluble.

LIQUAC and Ming and Russell model both show the smallest deviations for two data types and Extended UNIQUAC for one data type. On the other hand, LIQUAC shows the biggest deviation for two data types, and Ming and Russell model, Modified UNIFAC and LIFAC each for one data type. LIQUAC has parameters for most of the 31 electrolytes, whereas the other models have parameters for less than a half of the electrolytes. It should be noted that even if the number of applicable data sets for two models are equal, one of them may include solutions that are more difficult to fit, leading to larger average deviations even though the model in itself is not worse than the other.

4. Models selected for fitting parameters for atmospheric organic-electrolyte-water mixtures

Only one of the five models considered above has parameters for common atmospheric organics. The Ming and Russell model has parameters for e.g. dicarboxylic and hydroxy acids, but parameters are available only for four electrolytes. The organics described in the other four models are mainly organic solvents, such as alcohols, but these are not common in the atmosphere. Furthermore, some of the most common atmospheric ions are not included these models. For this reason, we decided to select models for fitting of new parameters for common atmospheric mixtures of organics and electrolytes.

When selecting models for fitting, not only deviations in water-electrolyte mixtures, but also the suitability of the model for water-organic-electrolyte solutions and organic-water solutions need to be considered. Also, the number of fitting parameters and need for experimental data, have great influence to the fitting. A selected model should be suitable for calculating both solvent and solute activities, and for solubility and vapour-liquid equilibrium calculations. All of the previous models should be suitable for organic-water and ternary solutions, but some principal differences exist. The biggest difference

is that LIFAC, Modified UNIFAC and Ming and Russell models are based on the group contribution method, and thus are predictive, whereas LIQUAC and Extended UNIQUAC are not predictive. Because the lack of experimental organic-electrolyte data for common atmospheric mixtures is the biggest problem in parameter fitting, we decided to select only predictive models for the fitting procedure.

4.1. Electrolytes and organic species selected for parameter fitting

Na^+, H^+, NH_4^+, NO_3^-, Cl^- and SO_4^{2-} are common atmospheric ions, and these ions were also selected for the parameter fitting. Because the models should be predictive, it is assumed that all electrolytes dissociate completely into ions, and the activities can then be calculated for each liquid phase ion. This means that ions are considered as functional groups which are not parts of other compounds, just like UNIFAC groups water (H_2O) and methanol (CH_3OH).

The organic fraction of atmospheric particles can be composed of several different organics, but organics with more than one polar functional group are very common. For this study, we selected organics which have at least two acid groups (COOH) and they may have hydroxyl groups (OH). Experimental data for fitting were available for dicarboxylic acids (oxalic, malonic, succinic, glutaric and adipic acid) and for citric, tartaric and malic acids which have both acid and hydroxyl groups. These non-electrolytes can be described with the functional groups H_2O, CH_n, OH and COOH.

4.2. The selected models

After considering deviations of model output from experimental data as well as the predictive capabilities of the models, we selected four models for fitting of new parameters. Here we outline the main features of the selected models briefly; detailed model equations and parameters are given in appendices. Two of the selected models are nearly the same as in the model comparison above, one model is a combination of two models considered above, and one model is actually the non-electrolyte part of

two of the models included in the model comparison. Because the model equations have already been published elsewhere, we are not going to examine the theory of the selected models, but refer instead to the original articles. However, we will examine the model accuracies, predictive capabilities and suitability for different calculations for atmospheric organic-electrolyte mixtures.

4.2.1. UNIFAC

The first model selected for fitting is the original UNIFAC (Fredenslund et al., 1975). New functional groups were added to the model for the selected ions. Because UNIFAC is a non-electrolyte model, it was not included in the model comparison in Sect. 3. However, UNIFAC is used as one component in many organic-electrolyte models, for example in LIFAC, Ming and Russell model and Modified UNIFAC, which uses slightly different Larsen's UNIFAC (Larsen et al., 1987). In these models UNIFAC has the biggest effect to the activities of non-electrolytes, and other terms have control over ion activities. Therefore, our model is not suitable for calculation of ion activities even if new parameters are fitted. However, this model is applied for calculating activities of organic species and water in both non-electrolyte and electrolyte solutions. From now on the term UNIFAC refers to our fitted UNIFAC, and when other UNIFAC models (e.g. the original UNIFAC) are used, it is stated in their names. The model equations and parameters are given in Appendix A.

4.2.2. LIFAC

Unlike LIQUAC (Li et al., 1994), LIFAC (Yan et al., 1999) is based on the group contribution method, and for this reason the second model selected for fitting is LIFAC. Model equations remain unaltered, but functional groups were added for the new organics and ions. LIFAC is a combination of a simple electrolyte model, original UNIFAC, and a term for non-electrolyte-ion interactions. The simple electrolyte model has no mixture parameters, but only pure component properties (e.g. density). LIFAC reduces to UNI-

FAC in non-electrolyte solutions, but all three parts are included in aqueous electrolyte solutions. The model equations and parameters are given in Appendix B. From now on the term LIFAC refers to our fitted LIFAC instead of the original model.

4.2.3. Ming and Russell

Ming and Russell (2002) model is one of the few activity coefficient models for atmospheric mixtures, and this model was selected as the third model for parameter fitting. Compared to the original model, one normalization term was dropped, and again functional groups were added for the new organics and ions. The reason for ignoring the normalization term is explained in Appendix C. This model is a combination of Clegg et al. (1992) electrolyte model and the original UNIFAC. The new model reduces to UNIFAC in non-electrolyte solutions, and to the Clegg et al. (1992) model in aqueous electrolyte solutions. The model equations and parameters are given in Appendix C. The original Ming and Russell (Clegg) model was the only model in the comparison of the published models in which the equilibrium for the incomplete dissociation of HSO_4^- ion was calculated, and this is done in the new model also. In addition to the selected ions, HSO_4^- ion was added to the model. In the other selected models, HSO_4^- ion is expected to dissociate into H^+ and SO_4^{2-} ions. From now on the term Ming and Russell model refers to our fitted model instead of the original one. When the original model is used, it is called the original Ming and Russell model.

4.2.4. Extended UNIFAC

The fourth model is a combination of Extended UNIQUAC (Thomsen et al., 1996) and Modified UNIFAC (Achard et al., 1994). Extended UNIQUAC is a combination of a simple electrolyte model and UNIQUAC (Abrams and Prausnitz, 1975) for non-electrolytes, but the UNIQUAC part makes this model non-predictive. Modified UNIFAC is a combination of a simple electrolyte model and Larsen's UNIFAC (Larsen et al., 1987), but the use of the hydrated mole fractions reduces model capabilities for concentrated elec-

trolyte solutions. By combining the electrolyte part of Extended UNIQUAC version of Iliuta et al. (2000) and the original UNIFAC, we have our new model. From now on, this new model is called Extended UNIFAC. Again, new groups for the selected organics and ions were added to the model. This model reduces to UNIFAC in non-electrolyte solutions. The model equations and parameters are given in Appendix D.

4.3. Parameter fitting

Because all new models reduces to UNIFAC in non-electrolyte solutions, we can use the same organic-water UNIFAC parameters. The original UNIFAC (Fredenslund et al., 1975) is quite accurate for simple organics like alcohols, but if one compound has two closely-spaced polar groups, the intramolecular interactions may cause large errors to the calculated activity coefficients. Improved results are obtained if new UNIFAC parameters are added or existing parameters are refitted for such compounds. Because the selected organics have at least two polar groups in relatively short carbon chain, we decided to refit all UNIFAC interaction parameters. However, pure component parameters for the solvent groups (surface and area parameters) are same as in the original UNIFAC.

The models treat electrolyte solutions differently, so the rest of the mixture parameters were fitted separately for each model using electrolyte-water and organic-electrolyte-water data. Most of the mixture parameters were fitted, but in addition to the non-electrolyte parameters, some of the interaction parameters were kept constant because of model considerations (e.g. UNIFAC interaction parameters for ions are zeros in LIFAC and Clegg et al. (1992) model parameters in the electrolyte part of the Ming and Russell model were not fitted). Furthermore, surface and area parameters for the ions were given constant values.

Model parameters were obtained by minimizing the sum of squared relative deviations of every data sets with the Nelder-Mead method (Nelder and Mead, 1965). The fitted non-electrolyte parameters, which are common for all new models, are presented in Appendix A and the remaining model-specific parameters are presented in Appen-

dices A–D for each model.

4.3.1. Experimental data

Experimental mixed electrolyte data can be found in the literature, however, we limit ourselves to considering single electrolytes. The single electrolyte data for $NaNO_3$, NaCl, Na_2SO_4, HNO_3, HCl, H_2SO_4, NH_4NO_3, NH_4Cl and $(NH_4)_2SO_4$, corresponding to the selected six ions, are the same as in the Table 2 except that water saturation activity data ($a_w^{sat}(T)$), which requires calculation of electrolyte solubility, was not included in the fitting. Instead of electrolyte solubility calculations during the minimization, we compared solubility products and products of ion activities calculated at given saturation molalities. Because our UNIFAC is for activity coefficients of non-electrolytes, electrolyte mean activity coefficient and salt solubility data were not used in the parameter fitting.

Sources and mixtures for non-electrolyte and organic-electrolyte data are given in Table 4. The non-electrolyte data in fitting includes water activities in aqueous single and binary organic solutions as a function of mole fractions at 298.15 K. The solubility and water activity data were used twice. In the first place, water and organic activities were calculated at given saturation concentration and temperature. The calculated water activities were then compared with experimental water activities, and organic activities were compared with those calculated from melting enthalpies. Because the melting enthalpies of oxalic and citric acids may be inaccurate, solubilities of these organics were not included in the fitting. However, these compounds were included in the model comparison, Sect. 4.4. The temperature interval for the solubility and activity data was from 273.15 K to 353.15 K, but sometimes data were available only at 298.15 K.

The organic-electrolyte data includes water activities as a function organic and electrolyte concentrations at constant temperature. The temperatures given in the different sources were between 293.15 K and 298.15 K. The last data type was water activity and concentrations in saturated water-organic-electrolyte solution at 297.15 K. Again,

instead of actually calculating solubilities, we calculated activities with given concentrations and then compared experimental and calculated water activities, organic activities and activities calculated from the melting enthalpies, and product of ion activities and solubility products. Oxalic acid saturation activities were not included in the fitting, because the melting enthalpy may be inaccurate.

4.4. Deviations from experimental data

In this section we examine the accuracies of the fitted models by comparing experimental and calculated water activities, solubilities and mean activity coefficients. Most of the experimental data was already used in the fitting, but also some new data sets are introduced. Deviations from experimental values were calculated for different data types as means of average absolute deviations of the data sets, similarly as in Sect. 3.2.

4.4.1. Deviations for non-electrolyte solutions

The data types for non-electrolyte solutions are water activity in organic-water solution as a function of organic mole fraction ($a_w(x_o)$), organic solubility as a function of temperature ($x_o^{sat}(T)$), water saturation activity as a function of solubility and temperature ($a_w(x_o^{sat}, T)$), water saturation activity as a function of temperature ($a_w^{sat}(T)$) and water activity in binary organic solution as a function of mole fractions ($a_w(x_{o1}, x_{o2})$). Mixtures and data sources for these data types are given in the Table 4.

Because all fitted models have the same UNIFAC parameters in non-electrolyte solutions, our model deviations are compared with those of three other UNIFAC-based non-electrolyte models. The three other models are the UNIFAC part of the original Ming and Russell model, the original UNIFAC, and UNIFAC fitted by Peng et al. (2001). Parameters for the original UNIFAC are from Hansen et al. (1991). Peng et al. (2001) fitted new UNIFAC interaction parameters for OH, H_2O and COOH, and the remaining parameters are the same as in the original UNIFAC. The experimental water activity data in Peng et al. (2001) fitting was also used in the fitting of our model parameters

(Table 4). Deviations for our refitted UNIFAC model and the three other UNIFAC-based models from the literature are given in Table 5.

The original UNIFAC shows the biggest deviations for all data types, but the original Ming and Russell model performs only slightly better. The present UNIFAC as well as that of Peng et al. (2001) have deviations that are usually less than half of the deviations of the original UNIFAC. Peng et al. (2001) UNIFAC has the smallest deviations for three data types and our model for two data types. Experimental and calculated water activities and organic solubilities for four organic species are presented in Fig. 1.

4.4.2. Deviations in electrolyte and organic-electrolyte solutions

Data types for the nine possible electrolyte-water mixtures are the same as in the comparison of the previously published models (Table 2). The only water-organic-electrolyte data type in this model comparison is water activity in organic-electrolyte-water mixtures as a function of mole fractions and salt molality ($a_w(x, m)$). The data sources and mixtures are given in the Table 4.

Model deviations and number of data sets are given in Table 6. Because the fitted UNIFAC is not suitable for ion activities, it was not used for the three data types in which ion activities are required. The number of experimental data sets is the same for all models except that LIFAC failed in $a_w^{sat}(T)$ calculation for $(NH_4)_2SO_4$. Originally, there were six data sets for solubility ($m^{sat}(T)$), but LIFAC failed to calculate the solubility of $(NH_4)_2SO_4$ at temperatures below 42°C (which also prevented the calculation of the $a_w^{sat}(T)$ of $(NH_4)_2SO_4$). On the other hand, the other models failed to calculate the solubility of NH_4NO_3 at temperatures above 40°C.

The fitted (and the original) Ming and Russell model reduce to the accurate Clegg et al. (1992) electrolyte model in the single electrolyte solutions, so it is no surprise that the model deviations are the smallest for four of the five single electrolyte data types. Extended UNIFAC has only a simple Debye-Hückel part for electrolytes, so considering the simplicity of the model, the deviations for water and ion activities are surprisingly small. UNIFAC shows the smallest deviation for one single electrolyte data

type and LIFAC shows the smallest deviation for the one organic-electrolyte data type. For comparison, deviation for the original Ming and Russell model in $a_w(x, m)$ data type is 0.07414, which is bigger than the deviations for the fitted models. Moreover, the original model has parameters only for nine of the eleven mixtures.

Even if both the original and the fitted Ming and Russell models are accurate for electrolyte solutions, their deviations in organic-electrolyte mixtures are quite big compared to the deviations of the other models. The main reason for this is that the original and the fitted Ming and Russell models have few fitting parameters for organic-electrolyte interactions.

The MR part of LIFAC is somewhat problematic for concentrated mixtures. In the MR part, the logarithms of activity coefficients are calculated as sums of terms which are proportional to the fitting parameters, ion molalities and ionic strength. Because these terms have quite large numerical values, a small change in the interaction parameters or molality can cause very big change to activity coefficients. This was the main reason for limiting electrolyte molalities to 30 mol/kg.

Experimental and calculated water activities, electrolyte solubilities and mean activity coefficients in single electrolyte solutions are presented in Figs. 2 and 3. Experimental and calculated water activities in aqueous organic-electrolyte solutions are presented in Fig. 4, where the original Ming and Russell model is included for comparison.

4.5. Growth factors for organic-salt particles

The accuracies and predictive capabilities of the fitted models are tested by calculating growth factors for organic-salt particles. For comparison, growth factors were also calculated with the original Ming and Russell model and ADDEM (Topping et al., 2005). Our UNIFAC and ADDEM (Aerosol Diameter Dependent Equilibrium Model) are not suitable for solubility calculations, so growth factors were calculated as if particles were dissolved at all relative humidities (supersaturated solutions). In ADDEM, the water content in mixed organic-electrolyte solutions is calculated from the water contents in organic-water and electrolyte-water solutions by using the ZSR (Stokes and Robinson,

1966) approach, in which organic-electrolyte interactions are ignored. The organic-water and electrolyte-water solutions are modelled with UNIFAC parameters from Peng et al. (2001) and Clegg et al. (1992) model respectively. Experimental and calculated growth factors for four different 100 nm organic-salt particles are plotted in Fig. 5. The dry particle composition is given as weight fractions.

Because there were few experimental data for electrolyte and organic activities, some deviations were expected to occur in the solubility calculations. The most difficult part in calculating of growth factors is the calculation of *MDRH*, where both the electrolyte and the organic species are saturated. It seems that most of the fitted models predict too low *MDRH*, but this depends very much on the organic-electrolyte mixture. At *MDRH*, when the liquid phase is formed, some fraction of the electrolyte or the organic species usually stays in the solid phase. When the relative humidity increases, condensing water dissolves part of the solid phase and this is balanced by further condensation of water resulting in faster growth compared to the pure liquid droplets. This kind of behaviour can be seen in the calculated growth factors. However, this behaviour is not seen in the experimental data, because variations in measurement data can be bigger than the effect of the dissolving solid phase to the growth factors. After the deliquescence, the models predict quite similar growth factors for three mixtures. The one exception is 80% glutaric acid – 20% NaCl mixture, where variations are quite large. However, the average of the predictions is quite close to the experimental values. In the case of 50% glutaric acid – 50% $(NH_4)_2SO_4$ mixture, all predicted growth factors are smaller than the experimental values at relative humidities above 85%.

Compared to the experimental accuracies of the other data types used in the model comparisons, there are much more uncertainties in the experimental growth factors. Cruz and Pandis (2000) data include few experimental values, and for example, in the case of 80% glutaric acid – 20% $(NH_4)_2SO_4$ particles, there seems to be a local maximum in the growth factors between 80 and 90% relative humidities, but this should not be possible. Data from Hämeri et al. (2002) is fairly smooth, however, Prenni et al. (2003) measured smaller growth factors for the same mixture at 303.15 K at *RH*'s

above 85%. These values would have been closer to the model predictions.

Other causes for the observed deviations between model predictions and experimental data are errors caused by approximation of solution densities and surface tensions. The density approximation can cause deviations to calculated droplet volume and to the Kelvin effect, where also the surface tension is needed. Because the diameters of the solid particles are 100 nm, the Kelvin term has only a small effect on the growth factors and *MDRH*. The volume approximation would be accurate for insoluble particles, and in the case of soluble particles, the effect of predicted water activity to liquid water content and through this to liquid volume is much bigger than the effect of solution density.

5. Summary and conclusions

In the first part of this paper we selected five different water-organic-electrolyte activity coefficient models from the literature. Only one of these models includes organics and electrolytes that are relevant in atmospheric modelling. Organics in the other models were organic solvents, such as alcohols. Furthermore, important atmospheric electrolytes (e.g. H^+) were missing form some of the models. Model deviations were calculated for different types of experimental single electrolyte data (Table 3) in order to assess model accuracy and suitability for atmospheric modelling.

In the second part of this paper four of the tested models were selected to be extended for organic-electrolyte mixtures relevant in the atmosphere (e.g. ammonia, sulphate, dicarboxylic acids and citric acid). The model selection was based on the model deviations and the predictive capabilities. Some of the selected models were modified if it was found useful, and both new and existing mixture parameters were fitted to experimental data. Model deviations were calculated for different non-electrolyte and electrolyte data types (Tables 5 and 6). For comparison, deviations were calculated also for some other models, which are suitable for these mixtures. Furthermore, model predictions and experimental data are presented in Figs. 1, 2, 3 and 4. Finally,

model predictions for growth factors of organic-electrolyte particles were compared to experimental data (Fig. 5).

Our results indicate that if experimental data is available, most of the predictive organic-electrolyte models can be successfully extended to atmospheric mixtures by fitting of new parameters. Because the organic fraction in atmospheric mixtures can be composed of many different compounds and there are few experimental data available for fitting, predictive models are especially useful in atmospheric modelling. For these reasons, only predictive models were selected for fitting.

The biggest problem in the fitting is lack of experimental data for organic activities in all solutions and electrolyte activities in organic-electrolyte solutions. The only organic activity data type was organic solubilities. Similarly, the only electrolyte activity data in ternary solutions was electrolyte solubilities. The biggest problem with this kind of data is that we have activities only at one single concentration. Another problem for organics was that melting enthalpies needed in solubility calculations were not found for all organics, but had to be estimated from other thermodynamic constants.

There are several accurate activity coefficient models for electrolytes and some of them are suitable for atmospheric mixtures. However, without fitting of new parameters, few of the non-electrolyte models are actually suitable for atmospheric organics. For example, the original UNIFAC showed much bigger deviations than our and the Peng et al. (2001) UNIFAC versions. There are very few models that are suitable for organic-electrolyte mixtures of atmospheric interest. We used one model (Ming and Russell, 2002) not fitted by us in the model comparison, and this model had bigger deviations for water activities in organic-electrolyte solution than our fitted models. There are some methods in which fitting parameters are not needed, for example ZSR (Stokes and Robinson, 1966) based ADDEM (Topping et al., 2005) and the Clegg et al. (2001) method, but deviations were not calculated for them. However, ADDEM was used in predicting growth factors for organic-electrolyte particles. Compared with the predictions of our models and experimental data, ADDEM overestimates growth factors in the systems studied.

If only water activities are needed, the simplest fitted model, i.e. UNIFAC, is reasonably accurate. For example, in the growth factor calculations, our UNIFAC performs better than most of the other models. In addition to good accuracy, UNIFAC is also very fast in computer simulations. In addition to the UNIFAC part, Extended UNIFAC has a simple electrolyte part, which makes this model suitable for calculating electrolyte activities in dilute single electrolyte solutions. However, because electrolyte activities were included in the fitting, deviations in organic-electrolyte solutions are bigger than those of UNIFAC. In addition to the simple electrolyte model and UNIFAC, LIFAC has a part describing non-electrolyte-ion interactions. This improves model accuracy in both electrolyte and organic-electrolyte solutions, where LIFAC has the smallest deviation. Because the fitted (and original) Ming and Russell model has few fitting parameters for the organic-electrolyte interactions, the model is close to an electrolyte model. For this reason, model deviations are the biggest in organic-electrolyte solutions, but small in single electrolyte solutions. Thanks to the accurate Clegg et al. (1992) electrolyte part, this model is suitable for complex mixtures of electrolytes.

There are very few activity coefficient models for ternary water-organic-electrolyte mixtures of atmospheric interest, and when predictions of our fitted models were compared to experimental data and to predictions of other models, our models performed well. Most of the published organic-electrolyte models are fitted for a few different mixtures, but our models had relatively large database in the fitting. In addition to species in the fitting, our predictive models can be used to predict activities for other similar compounds as well. Because there are few experimental data and other models for these mixtures, the fitted models can be very useful.

Appendix A: UNIFAC

The first model is the same as the original UNIFAC (Fredenslund et al., 1975), except that we introduced new functional groups for ions and organics. This has no effect to the model equations, but new model parameters are needed. Because this new

model is not used for ion activities, the reference states are the same as in the original UNIFAC. However, in the other three fitted models in which ion activities are calculated, ions have different reference states than non-electrolytes.

UNIFAC (Fredenslund et al., 1975) is so called group contribution method, where chemical species are constructed from functional groups. This reduces the number of fitting parameters and makes the model predictive. Activity coefficient for species i is calculated as a sum of combinatorial and residual contributions

$$\ln \gamma_i = \ln \gamma_i^{\mathrm{C}} + \ln \gamma_i^{\mathrm{R}} \tag{A1}$$

The combinatorial contribution is calculated with the equation

$$\ln \gamma_i^{\mathrm{C}} = 1 - V_i + \ln V_i - \frac{z}{2} q_i \left(1 - \frac{V_i}{F_i} + \ln\left(\frac{V_i}{F_i}\right)\right) \tag{A2}$$

where $V_i = \frac{r_i}{\sum_k r_k x_k}$, $F_i = \frac{q_i}{\sum_k q_k x_k}$ and $z=10$. x_i is mole fraction, $q_i = \sum_k v_k^{(i)} Q_k$ is area parameter, $r_i = \sum_k v_k^{(i)} R_k$ is volume parameter and $v_k^{(i)}$ is the number of functional group k in species i. The only parameters in the combinatorial part are the pure component area (Q_k) and volume (R_k) parameters for sub groups.

The residual contribution is calculated with the equation

$$\ln \gamma_i^{\mathrm{R}} = \sum_k v_k^{(i)} \left[\ln \Gamma_k - \ln \Gamma_k^{(ref,i)}\right] \tag{A3}$$

where Γ_k describes the effect of sub group k on the activity coefficient of component i. The latter term in the sum describes the effect of sub group k in the reference state of component i, making the activity coefficient unity at the reference state. The reference state for non-electrolytes is liquid of pure component i. In our electrolyte models, the reference state for ion i is infinitely dilute aqueous solution. The equation for Γ_k is

$$\ln \Gamma_k = Q_k \left[1 - \ln\left(\sum_m \frac{X_m Q_m}{\sum_n X_n Q_n} \Psi_{mk}\right) - \sum_m \left(\frac{X_m Q_m \Psi_{km}}{\sum_n X_n Q_n \Psi_{nm}}\right)\right] \tag{A4}$$

where $X_k = \frac{\sum_j \nu_k^{(j)} x_j}{\sum_j \sum_m \nu_m^{(j)} x_j}$ is mole fraction of group k in the mixture, $\Psi_{km} = \exp\left(-\frac{a_{km}}{T}\right)$ and a_{km} is the group interaction parameter for main groups k and m. Most of the functional groups (main groups) have sub groups which have the same interaction parameters, but different Q_k and R_k. For example the main group CH_n has sub groups CH_3, CH_2, CH and C.

The group volume and surface area parameters can be calculated from the molecular sizes of the functional groups. This is difficult for ions because for example the H^+ ion is very small compared to the sizes of organic groups. These parameters could also be fitted, but this can lead to very small or high values. In the new UNIFAC model, the group volume and surface area parameters for non-electrolytes are the same as in the original UNIFAC, and parameters for ions were taken from the literature. Group volume and surface area parameters for all functional groups are given in Table 7. Fitted interaction parameters for non-electrolytes are in Table 8. These interaction, surface area and volume parameters are the same for all fitted models where parameters are needed. Ion-organic interaction parameters were also fitted for this model and are given in Table 9. Ion-ion interaction parameters were set to zero.

Appendix B: LIFAC

The second model is the same as the original LIFAC (Yan et al., 1999), with the addition of new functional groups for electrolytes and organics. The main and the sub groups are same as in our UNIFAC model. Activity coefficients in LIFAC (Yan et al., 1999) are calculated as a sum of three contributions: long range (LR), middle range (MR) and short range (SR).

$$\ln \gamma_i = \ln \gamma_i^{LR} + \ln \gamma_i^{MR} + \ln \gamma_i^{SR} \tag{B1}$$

The LR contribution is calculated with a Debye-Hückel term, the MR contribution is calculated with virial equation and the SR contribution is calculated with UNIFAC. When

electrolyte concentration is zero, the LR and MR contributions for all species are zeros and the model reduces to UNIFAC.

The long range contribution is calculated with Debye-Hückel theory. This part has no fitting parameters, but only pure solvent properties. The solvent is a mixture of water and organic compound, and its properties are calculated from pure component properties weighted with salt free volume fraction $\phi'_s = \frac{x'_s \frac{M_s}{\rho_s}}{\sum_{i \neq salt} x'_i \frac{M_i}{\rho_i}}$, where M_s (kg/mol) is molecular weight, ρ_s (kg/m^3) is density and x'_s is salt free mole fraction of pure solvent s. Solvent (s) and ion (i) activity coefficients are calculated with the equations

$$\ln \gamma_s^{\mathrm{LR}} = \frac{2AM_s\rho}{b^3\rho_s}\left(1 + b\sqrt{I} - \frac{1}{1 + b\sqrt{I}} - 2\ln\left(1 + b\sqrt{I}\right)\right) \tag{B2}$$

$$\ln \gamma_i^{\mathrm{LR}} = \frac{-z_i^2 A\sqrt{I}}{1 + b\sqrt{I}} \tag{B3}$$

where $\rho = \sum_s \phi'_s \rho_s$ is density of the solvent mixture, z_i is charge magnitude of ion i and $I = 0.5\sum_i m_i z_i^2$ (mol/kg) is ionic strength in the molality (m) scale. The Debye-Hückel parameters A ($\sqrt{\mathrm{kg/mol}}$) and b ($\sqrt{\mathrm{kg/mol}}$) are calculated with equations

$$A = e_0^3 \frac{(2\pi N_A \rho)^{1/2}}{(DkT)^{3/2}} \tag{B4}$$

$$b = a\sqrt{\frac{8\pi e_0^2 N_A \rho}{DkT}} \tag{B5}$$

where e_0 (C) is elementary charge, N_A (1/mol) is Avogadro's constant, k (J/K) is Boltzmann constant, T (K) is temperature and a (m) is the closest approach parameter, here $a = 10^{-10}$ m. Dielectric constant is calculated with equation $D = 4\pi\epsilon_0\epsilon_r$, where ϵ_0 (C^2/(J m)) is permittivity of vacuum and ϵ_r is relative permittivity. In the original model,

D is dielectric constant of solvent mixture, but because dielectric constants were not available for all organics, we use that of pure water instead.

The middle range contribution is calculated using functional groups, which are the same as in the UNIFAC part. The equations for solvent group k and ion j are

$$\ln \gamma_k^{\mathrm{MR}} = \sum_i B_{ki} m_i - \frac{M_k \sum_k \sum_i v_k^{(i)} x_i'}{M} \sum_k \sum_i \left[B_{ki} + I B_{ki}'\right] x_k' m_i - M_k \sum_c \sum_a \left[B_{ca} + I B_{ca}'\right] m_c m_a \tag{B6}$$

$$\ln \gamma_j^{\mathrm{MR}} = \frac{1}{M} \sum_k B_{jk} x_k' + \frac{z_j^2}{2M} \sum_k \sum_i B_{ki} x_k' m_i + \sum_i B_{ji} m_i + \frac{z_j^2}{2} \sum_c \sum_a B_{ca}' m_c m_a \tag{B7}$$

where B_{jk} (kg/mol) is interaction coefficient for groups j and k, B_{jk}' (kg^2/mol^2) is derivative of B_{jk} with respect to ionic strength, $M = \sum_s x_s' M_s$ is molecular weight of solvent mixture and M_k is molecular weight of solvent group k, which were calculated from atomic weights. Subscripts k, i, c and a denote solvent group, any ion, cation and anion respectively. Symmetric ($B_{jk} = B_{kj}$) interaction coefficients are functions of ionic strength

$$B_{ca}(I) = b_{ca} + c_{ca} \exp(-\sqrt{I} + 0.13 I) \tag{B8}$$

$$B_{ki}(I) = b_{ki} + c_{ki} \exp(-1.2\sqrt{I} + 0.13 I) \tag{B9}$$

where b_{jk} and c_{jk} are interaction parameters for main groups j and k. Ion activity coefficients must be normalized to infinite dilution reference state by subtracting the first term in Eq. (B7) calculated with correct reference state concentrations. The MR activity coefficient of solvent s is calculated as a sum of group activity coefficients

$$\ln \gamma_s^{\mathrm{MR}} = \sum_k v_k^{(s)} \ln \gamma_k^{MR} \tag{B10}$$

Some of the MR interaction parameters could have been taken directly from the original model, but because two of the six ions were new, all parameters were fitted. The fitted MR ion-solvent group and cation-anion parameters are given in Table 10.

The short range contribution is calculated with UNIFAC (see Appendix A), but now the infinite dilution reference state is selected for the ions. Because UNIFAC equations give activity coefficients in mole fraction scale, ion activity coefficients are converted to molality scale with equation

$$\ln \gamma_i^{(m)} = \ln \gamma_i^{(x)} - \ln\left(M_r/M + M_r \sum_i m_i\right) \tag{B11}$$

where M_r is molecular weight of the reference state solvent and M is molecular weight of the solvent mixture. The SR part UNIFAC parameters are given in the Tables 7 and 8. All non-electrolyte-ion and ion-ion UNIFAC interaction parameters are set to zeros.

Appendix C: Ming and Russell model

The third model is nearly the same as the original Ming and Russell (2002) model, but one normalization term was dropped and new groups for ions and organics were added. Activity coefficients are calculated as a sum of ion-water interactions (IW) and organic-water/organic-ion interactions (OW/OI).

$$\ln \gamma_i = \ln \gamma_i^{\mathrm{IW}} + \ln \gamma_i^{\mathrm{OW/OI}} \tag{C1}$$

Ion-water interactions are calculated with the Clegg et al. (1992) model and organic-water/organic-ion interactions are calculated with UNIFAC. If the organic concentration is zero, the new model is reduced to Clegg et al. (1992) model and if the ion concentration is zero the model is reduced to UNIFAC.

Ion-water interaction are calculated with equations from Clegg et al. (1992) as a sum of LR and SR contributions

$$\ln \gamma_i^{\mathrm{IW}} = \ln \gamma_i^{\mathrm{LR}} + \ln \gamma_i^{\mathrm{SR}} \tag{C2}$$

Ming and Russell calculated the LR and SR activity coefficients using inorganic-only mole fractions x_i^* (LR* and SR*), where organics (subscript o) are ignored. The activity coefficients were then normalized to the solution mole fractions with the equation

$$\ln \gamma_i^{\mathrm{IW}} = \ln \gamma_i^{\mathrm{LR*}} + \ln \gamma_i^{\mathrm{SR*}} - 2 \ln \left(1 - \sum_o x_o \right) \tag{C3}$$

Even if we use these inorganic-only mole fractions in our modified version, the correction term is ignored. The reason for this is that for non-electrolyte solutions (x_w^*=1 and x_{ion}^*=0) the original model does not reduce to UNIFAC, because the correction term is not zero although it is based on the difference in mole fractions calculated with and without organics. In fact, Ming and Russell (2002) had an error in their computer code which fortuitously ensured that the model reduced to the UNIFAC in non-electrolyte solutions (Y. Ming, personal communication, 2003).

All mole fractions in the IW part are inorganic-only mole fractions, so for clarity, the superscript * is dropped from mole fractions in following model equations. The Clegg et al. (1992) model equations are given in condensed form and all terms that include ternary parameters were ignored. Thus, these equations are valid only if ternary parameters are zeros as is the case e.g. in aqueous sulphuric acid system and in all binary systems. The complete model equations are given in Clegg et al. (1992).

The long range contribution is a sum of Debye-Hückel (DH) contribution and a higher order electrostatic (HOE) contribution to the Debye-Hückel expression.

$$\ln \gamma_i^{\mathrm{LR}} = \ln \gamma_i^{\mathrm{DH}} + \ln \gamma_i^{\mathrm{HOE}} \tag{C4}$$

The DH contribution for water w and ion i are calculated with equations

$$\ln \gamma_w^{\mathrm{DH}} = \frac{2A_x I_x^{3/2}}{1+\rho\sqrt{I_x}} - \sum_c \sum_a x_c x_a \left[B_{ca} \exp\left(-\alpha_{ca}\sqrt{I_x}\right) + B_{ca}^1 \exp\left(-\alpha_{ca}^1\sqrt{I_x}\right)\right] \quad \text{(C5)}$$

$$\begin{aligned}\ln \gamma_i^{\mathrm{DH}} = {} & \frac{2A_x I_x^{3/2}}{1+\rho\sqrt{I_x}} - \sum_c \sum_a x_c x_a \left[B_{ca} \exp\left(-\alpha_{ca}\sqrt{I_x}\right) + B_{ca}^1 \exp\left(-\alpha_{ca}^1\sqrt{I_x}\right)\right] - \\ & z_i^2 A_x \left(\frac{2}{\rho} \ln\left(1+\rho\sqrt{I_x}\right) + \frac{\sqrt{I_x}}{1+\rho\sqrt{I_x}} \right) + \\ & \sum_{j\neq i} x_j \left(B_{ij}\, g\left(\alpha_{ij}\sqrt{I_x}\right) + B_{ij}^1\, g\left(\alpha_{ij}^1\sqrt{I_x}\right)\right) - \\ & \frac{z_i^2}{2I_x} \sum_c \sum_a x_c x_a \left[B_{ca} \left[g\left(\alpha_{ca}\sqrt{I_x}\right) - \exp\left(-\alpha_{ca}\sqrt{I_x}\right)\right] + \right. \\ & \left. B_{ca}^1 \left[g\left(\alpha_{ca}^1\sqrt{I_x}\right) - \exp\left(-\alpha_{ca}^1\sqrt{I_x}\right)\right]\right] \end{aligned} \quad \text{(C6)}$$

where constant ρ=13.0 and the temperature dependent mole fraction scale Debye-Hückel parameter A_x has a value 2.917 at 298.15 K. I_x is ionic strength in mole fraction scale and function $g(x)=\frac{2(1-(1+x)\exp(-x))}{x^2}$. Most of the symmetric ($B_{ij}=B_{ji}$) parameters are zeros: $B_{cc}=B_{aa}=B_{cc}^1=B_{aa}^1=0$ and $\alpha_{cc}=\alpha_{aa}=\alpha_{cc}^1=\alpha_{aa}^1=0$. The remaining model parameters for DH part are B_{ca}, B_{ca}^1, α_{ca} and α_{ca}^1.

The HOE contribution for water w and ion i is needed if we have more than two ions (here only in the case of aqueous sulphuric acid). The equations are

$$\ln \gamma_w^{\mathrm{HOE}} = -2 \sum_c \sum_{<c'} x_c x_{c'} \left(\vartheta_{cc'} + I_x \vartheta'_{cc'}\right) - 2 \sum_a \sum_{<a'} x_a x_{a'} \left(\vartheta_{aa'} + I_x \vartheta'_{aa'}\right) \quad \text{(C7)}$$

$$\ln \gamma_i^{\mathrm{HOE}} = 2\sum_{j\neq i} x_j \left[\vartheta_{ij} - x_i\left(\vartheta_{ij} + \vartheta'_{ij}\left(I_x - \frac{z_i^2}{2}\right)\right)\right] - 2\sum_{c\neq i}\sum_{<c'\neq i} x_c x_{c'}\left(\vartheta_{cc'} + \vartheta'_{cc'}\left(I_x - \frac{z_i^2}{2}\right)\right) - 2\sum_{a\neq i}\sum_{<a'\neq i} x_a x_{a'}\left(\vartheta_{aa'} + \vartheta'_{aa'}\left(I_x - \frac{z_i^2}{2}\right)\right) \quad \text{(C8)}$$

where subscript j includes all cations if i is a cation or all anions if i is a anion, and

$$\vartheta_{ij} = \frac{z_i z_j}{4I_x}\left[J(x_{ij}) - \frac{1}{2}J(x_{ii}) - \frac{1}{2}J(x_{jj})\right]$$

$$J(x_{ij}) = \frac{x_{ij}}{4 + C_1 x_{ij}^{C_2}\exp(C_3 x_{ij}^{C_4})}$$

$$x_{ij} = 6 z_i z_j A_x \sqrt{I_x}$$

The constants are C_1=4.581, C_2=−0.7237, C_3=−0.012 and C_4=0.528. Derivatives are $\vartheta'_{ij}=\frac{\partial\vartheta_{ij}}{\partial I_x}$ and $J'(x_{ij})=\frac{\partial J(x_{ij})}{\partial x_{ij}}$. The HOE part has no fitting parameters.

The short range contributions for water w, cation C and anion A are calculated with equations

$$\ln \gamma_w^{SR} = \sum_c\sum_a \left(\frac{1}{F}E_c E_a \frac{z_c + z_a}{z_c z_a}(1 - x_w)W_{ca} + x_c x_a \frac{(z_c + z_a)^2}{z_c z_a}(1 - 2x_w)U_{ca} + 4x_c x_a x_w(2 - 3x_w)V_{ca}\right) \quad \text{(C9)}$$

$$\ln \gamma_C^{SR} = \sum_a\sum_{c\neq C} E_a\left[\frac{z_C}{2}E_c\frac{z_c + z_a}{z_c z_a}W_{ca}\right] - \sum_a\sum_c\left[x_w E_c E_a\left(\frac{z_C}{2} + \frac{1}{F}\right)\frac{z_c + z_a}{z_c z_a}W_{ca} +\right.$$

$$2x_w x_c x_a \frac{(z_c+z_a)^2}{z_c z_a} U_{ca} + 12 x_w^2 x_c x_a V_{ca}\Bigg] + \sum_a \Bigg[x_w E_a \frac{z_C+z_a}{z_a} W_{Ca} +$$

$$x_w x_a \frac{(z_C+z_a)^2}{z_C z_a} U_{Ca} + 4 x_w^2 x_a V_{Ca} - E_a\left(1-\frac{E_C}{2}\right)\frac{z_C+z_a}{z_a} W_{Ca}\Bigg] \quad \text{(C10)}$$

$$\ln \gamma_A^{SR} = \sum_c \sum_{a\neq A} E_c \left[\frac{z_A}{2} E_a \frac{z_c+z_a}{z_c z_a} W_{ca}\right] - \sum_c \sum_a \Bigg[x_w E_c E_a \left(\frac{z_A}{2}+\frac{1}{F}\right)\frac{z_c+z_a}{z_c z_a} W_{ca} +$$

$$2x_w x_c x_a \frac{(z_c+z_a)^2}{z_c z_a} U_{ca} + 12 x_w^2 x_c x_a V_{ca}\Bigg] + \sum_c \Bigg[x_w E_c \frac{z_c+z_A}{z_c} W_{cA} +$$

$$x_w x_c \frac{(z_c+z_A)^2}{z_c z_A} U_{cA} + 4 x_w^2 x_c V_{cA} - E_c\left(1-\frac{E_A}{2}\right)\frac{z_c+z_A}{z_c} W_{cA}\Bigg] \quad \text{(C11)}$$

where lower case letters refer to any other cation and anion, $E_c = \frac{x_c z_c}{\sum_c x_c z_c}$, $E_a = \frac{x_a z_a}{\sum_a x_a z_a}$. Fitting parameters for the SR part are symmetric W_{ca}, U_{ca} and V_{ca}.

These equations give ion activity coefficients in mole fraction scale, so these must be converted to molality scale with Eq. (B11). Parameters for IW part were not fitted, but these are from Clegg et al. (1992); Clegg and Brimblecombe (1995); Clegg et al. (1998a). Parameters are presented in Table 11.

OW/OI interactions are calculated with UNIFAC. Also here, the ion activity coefficients are normalized to infinite dilution reference state and from mole fraction scale to molality scale with Eq. (B11). In the original Ming and Russell model, surface area and volume parameters for ions were the same as for water, and all ion-water and ion-ion interaction parameters were zeros. For this reason, the OW/OI contribution for water and electrolytes is zero if the organic fraction is zero. The same values were selected to our modified model. Organic-water UNIFAC parameters are the same as in the other fitted models (Tables 7 and 8). Organic-ion interaction parameters were fitted and these are given in Table 12.

Appendix D: Extended UNIFAC

The fourth model, called Extended UNIFAC, is a modification of Extended UNIQUAC (Thomsen et al., 1996) and Modified UNIFAC (Achard et al., 1994). Activity coefficients are calculated as a sum of LR and SR contributions

$$\ln\gamma_i = \ln\gamma_i^{\mathrm{LR}} + \ln\gamma_i^{\mathrm{SR}} \tag{D1}$$

The LR part is the same as the Debye-Hückel part in Iliuta et al. (2000) version of Extended UNIQUAC, and the SR part is the UNIFAC with new functional groups for ions and organics.

The LR contribution is always zero for organics, but molality scale equations for water (w) and ion (i) activity coefficients from Iliuta et al. (2000) are

$$\ln\gamma_w^{\mathrm{LR}} = \frac{2AM_w}{b^3}\left(1 + b\sqrt{I} - \frac{1}{1+b\sqrt{I}} - 2\ln\left(1+b\sqrt{I}\right)\right) \tag{D2}$$

$$\ln\gamma_i^{\mathrm{LR}} = \frac{-z_i^2 A\sqrt{I}}{1+b\sqrt{I}} \tag{D3}$$

In the Extended UNIQUAC, the constant b has the value 1.50 $\mathrm{kg}^{1/2}\,\mathrm{mol}^{-1/2}$ and molality scale Debye-Hückel parameter A ($\mathrm{kg}^{1/2}\,\mathrm{mol}^{-1/2}$) is given as a function of temperature based on the temperature dependence of density and dielectric constant of pure water:

$$A = 1.131 + 1.335\cdot 10^{-3}(T/\mathrm{K} - 273.15) + 1.164\cdot 10^{-5}(T/\mathrm{K} - 273.15)^2 \tag{D4}$$

This equation is valid at temperatures 273.15 K$<T<$383.15 K. These constants are also used in the Extended UNIFAC.

The SR part is the same as in the three other models. Again, ion activity coefficients are normalized to infinite dilution reference state and to molality scale. Organic-water interaction parameters (Table 7), and surface and area parameters for all species (Table 8) are the same as in the other models. Ion-solvent and cation-anion interaction

parameters for Extended UNIFAC are given in Table 13. Because we had only single electrolyte data, cation-cation and anion-anion interaction parameters were not fitted.

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Table 1. Properties of the organic acids. Melting enthalpy ΔH_m (kJ/mol) and melting temperature T_m (K) for malonic acid are from Hansen and Beyer (2004), melting enthalpies were calculated for oxalic and citric acids, and for other organics these values are from NIST Chemistry webBook (http://webbook.nist.gov). Molecular weight M (g/mol) and density ρ (kg/m^3) for adipic acid are from Hori et al. (2003) and for other organics these are from Peng et al. (2001).

Acid	Formula	ΔH_m	T_m	M	ρ
oxalic	$C_2H_2O_4$	26.87	464.45	90.04	1900
malonic	$C_3H_4O_4$	24.94	407.46	104.06	1630
succinic	$C_4H_6O_4$	33.12	457.0	118.09	1552
glutaric	$C_5H_8O_4$	23.36	371.0	132.12	1429
adipic	$C_6H_{10}O_4$	34.85	426.4	146.14	1360
citric	$C_6H_8O_7$	37.52	427.15	192.12	1665
tartaric	$C_4H_6O_6$	32.30	445.1	150.09	1759.8
malic	$C_4H_6O_5$	33.52	402.0	134.09	1609

Table 2. Sources for experimental single electrolyte data. Water activity data for HNO_3 was generated from fitting of Clegg and Brimblecombe (1990). Water activities for electrolytes from Hamer and Wu (1972) were calculated from osmotic coefficients.

Mean activity coefficient as a function of salt molality, $\gamma_{\pm}(m)$	
Clegg et al. (1996): $(NH_4)_2SO_4$	Rard and Clegg (1997): $CaCl_2$
Hamer and Wu (1972): HNO_3, HCl, HN_4NO_3, NH_4Cl, $NaNO_3$, NaCl, NaBr, NaOH, NaI, LiCl, KCl, KBr, LiBr, KNO_3, KI, KF, $LiNO_3$, LiI	Robinson and Stokes (1959): H_2SO_4, Na_2SO_4, Li_2SO_4, K_2SO_4, $Ca(NO_3)_2$, $CaBr_2$, $ZnSO_4$, BaI_2, $BaBr_2$, $ZnCl_2$, $CuCl_2$
Water activity as a function of molality, $a_w(m)$	
Albright et al. (2000): $ZnSO_4$	El Guendouzi et al. (2003): $(NH_4)_2SO_4$, Na_2SO_4, Li_2SO_4, K_2SO_4
Clegg and Brimblecombe (1990): HNO3	Hamer and Wu (1972): NH_4NO_3, NaBr, NaOH, NaI, KBr, LiBr, KI, KF, $LiNO_3$, LiI
Correa et al. (1997): $NaNO_3$, KNO_3	Rard and Clegg (1997): $CaCl_2$
El Guendouzi and Dinane (2000): LiCl	
El Guendouzi et al. (2001): HCl, NH_4Cl	
Robinson and Stokes (1959): H_2SO_4, NaCl, KCl	
Salt solubility as a function of temperature, $m^{sat}(T)$ (mol/kg)	
Apelblat and Korin (1998a): $NaNO_3$, NaCl, NaBr	ChemDAT*: $(NH_4)_2SO_4$, NH_4NO_3, NH_4Cl, Na_2SO_4, $NaNO_3$, NaCl, K_2SO_4, NaBr, NaOH, NaI, KCl, LiBr, KNO_3, KI, $LiNO_3$, $ZnCl_2$, LiI
Apelblat and Korin (1998b): Li_2SO_4, KNO_3, KI	
Apelblat and Korin (2002): Na_2SO_4	
Pinho and Macedo (2002): KBr	
Water activity in saturated solution as a function of saturation molality and temperature, $a_w(m^{sat}, T)$	
Apelblat and Korin (1998a): $NaNO_3$, NaCl, NaBr	Apelblat and Korin (1998b): Li_2SO_4, KNO_3, KI
Water activity in saturated solution as a function of temperature, $a_w^{sat}(T)$	
Apelblat and Korin (1998a): $NaNO_3$, NaCl, NaBr	Greenspan (1977): $(NH_4)_2SO_4$, NH_4Cl, $NaNO_3$, NaCl, K_2SO_4, NaBr, NaOH, NaI, LiCl, KCl, KBr, LiBr, $CaBr_2$, KNO_3, KI, KF, LiI
Apelblat and Korin (1998b): Li_2SO_4, KNO_3, KI	
Lightstone et al. (2000): NH_4NO_3	
Seinfeld and Pandis (1998): Na_2SO_4	

*ChemDAT – The Merck Chemical Databases (http://chemdat.merck.de/mda/int_en/labtool/index.html)

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Table 3. Experimental single electrolyte data types, number of data sets (N) and deviations for each published activity coefficient model. The unit for deviations in solubility data ($m^{sat}(T)$) is mol/kg.

	LIQUAC		LIFAC		Ext. UNIQUAC		Mod. UNIFAC		Ming and Russell	
Data type	N	deviation	N	deviation	N	deviation	N	deviation	N	deviation
$\gamma_{\pm}(m)$	26	0.3884	15	0.5768	13	0.5827	15	1.281	12	1.373
$a_w(m)$	20	0.001941	12	0.005822	13	0.01162	16	0.02283	11	0.002188
$m^{sat}(T)$	12	7.369	9	4.498	9	3.758	12	5.089	7	0.8100
$a_w(m^{sat}, T)$	5	0.06602	5	0.09132	3	0.01689	5	0.08204	2	0.01801
$a_w^{sat}(T)$	14	0.05173	8	0.04705	9	0.02077	15	0.04780	8	0.01181

Table 4. Sources for experimental non-electrolyte and organic-electrolyte data used in the parameter fitting and model comparison. In addition to data where solubilities of oxalic and citric acids are needed, mixtures marked with asterisks were not used in the parameter fitting.

Water activity in organic-water solution as a function of mole fraction	
Maffia and Meirelles (2001)	oxalic, malonic, succinic, citric, tartaric and malic acids
Peng et al. (2001)	oxalic*, malonic*, succinic*, glutaric, citric*, tartaric* and malic* acids
Apelblat et al. (1995a)	citric acid*
Apelblat et al. (1995b)	citric* and tartaric* acids
Levien (1955)	citric acid*
Water activity in binary organic solutions as a function of mole fraction	
Marcolli et al. (2004)	malonic acid + malic acid
Maffia and Meirelles (2001)	citric acid + malic acid
Water activity and organic solubility as a function of temperature	
Apelblat et al. (1995a)	citric, tartaric and malic acids
Marcolli et al. (2004)	malonic, succinic, glutaric and adipic acids
Peng et al. (2001)	oxalic acid
Srinivasakannan et al. (2002)	oxalic acid (solubility only)
Water activity in aqueous organic-electrolyte mixture as a function of concentrations	
Choi and Chan (2002)	$(NH_4)_2SO_4$ + malonic/succinic/glutaric/citric acid
	NaCl + malonic/succinic/glutaric/citric acid
Lightstone et al. (2000)	NH_4NO_3 + succinic acid
Schunk and Maurer (2004)	Na_2SO_4/$NaNO_3$/NaCl + citric acid
Water activity and saturation concentrations in saturated water-organic-electrolyte mixture	
Brooks et al. (2002)	$(NH_4)_2SO_4$ + oxalic/malonic/succinic/glutaric/adipic/malic acid

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Table 5. Experimental non-electrolyte data types, number of mixtures (N) and deviations for our fitted UNIFAC and three other UNIFAC models. In addition to experimental data in the Table 4, data from Apelblat et al. (1995a,b); Levien (1955); Srinivasakannan et al. (2002) was included.

Data type	N	refitted UNIFAC	orig. Ming and Russell	orig. UNIFAC	Peng et al. (2001)
$a_w(x_o)$	7	0.02918	0.06483	0.07775	0.02265
$x_o^{sat}(T)$	8	0.01618	0.06994	0.08314	0.03060
$a_w(x_o^{sat}, T)$	8	0.03562	0.03710	0.0452	0.01657
$a_w^{sat}(T)$	8	0.04688	0.1043	0.1137	0.05426
$a_w(x_{o1}, x_{o2})$	2	0.02011	0.06268	0.07961	0.01548

Table 6. Experimental single electrolyte and organic-electrolyte data types, number of mixtures (N) and deviations for the fitted models. Sources for experimental single electrolyte data are given in the Table 2 and sources for organic-electrolyte data are given in the Table 4. The unit for deviations in solubility data ($m^{sat}(T)$) is mol/kg.

Data type	N	UNIFAC	LIFAC	Ming and Russell	Ext. UNIFAC
$\gamma_{\pm}(m)$	9		0.1279	0.07168	0.5023
$a_w(m)$	9	0.02509	0.009838	0.001104	0.01390
$m^{sat}(T)$	5		5.323	1.019	1.512
$a_w(m^{sat}, T)$	2	0.01152	0.01582	0.01801	0.02421
$a_w^{sat}(T)$	6		0.01744	0.01219	0.02521
$a_w(x, m)$	11	0.02824	0.01798	0.05555	0.04546

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Table 7. UNIFAC group volume R_k and surface area Q_k parameters for the modelled electrolytes and non-electrolytes. Parameters for the non-electrolytes are the same as in the original UNIFAC from Hansen et al. (1991). Parameters for Na^+, NH_4^+, Cl^- and NO_3^- are from Yan et al. (1999). Parameters for H^+ are the same as for most cations in Yan et al. (1999) and parameters for SO_4^{2-} are the same as for NO_3^-.

Main group	Sub group	R_k	Q_k
H_2O	H_2O	0.9200	1.400
CH_n	CH_2	0.6744	0.540
CH_n	CH	0.4469	0.228
CH_n	C	0.2195	0.000
OH	OH	1.0000	1.200
$COOH$	$COOH$	1.3010	1.224
Na^+	Na^+	3.0000	3.000
NH_4^+	NH_4^+	3.0000	3.000
H^+	H^+	1.0000	1.000
Cl^-	Cl^-	0.9861	0.992
NO_3^-	NO_3^-	1.6400	1.600
SO_4^{2-}	SO_4^{2-}	1.6400	1.600

Table 8. Fitted UNIFAC interaction parameters a_{ij} (K) for the modelled non-electrolytes.

i	j	a_{ij}	a_{ji}
H_2O	CH_n	170.22	2650.8
H_2O	OH	−1.3932	−407.50
H_2O	COOH	−437.73	271.04
CH_n	OH	143.48	19.236
CH_n	COOH	−150.91	2693.3
OH	COOH	−492.09	238.13

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Table 9. Fitted UNIFAC non-electrolyte-ion interaction parameters a_{ij} (K). Some parameters are zeros because of lack of experimental data. Water-organic interaction parameters are in the Table 8, and all ion-ion interaction parameters were set to zero.

i	j	a_{ij}	a_{ji}	i	j	a_{ij}	a_{ji}
H_2O	H^+	−1998	−735.4	OH	H^+	0	0
H_2O	Na^+	48.27	−287.7	OH	Na^+	1.804	−12.22
H_2O	NH_4^+	−558.6	2817	OH	NH_4^+	−20.09	79.73
H_2O	Cl^-	48.31	132.0	OH	Cl^-	0.2297	−423.1
H_2O	SO_4^{2-}	−1818	1620	OH	SO_4^{2-}	−1877	4062
H_2O	NO_3^-	367.6	1368	OH	NO_3^-	39.43	189.8
CH_n	H^+	0	0	COOH	H^+	0	0
CH_n	Na^+	−473.0	11.07	COOH	Na^+	1440	636
CH_n	NH_4^+	−2.452	4478	COOH	NH_4^+	−632.9	−167.6
CH_n	Cl^-	−640.5	39.66	COOH	Cl^-	2135	−234.9
CH_n	SO_4^{2-}	640.5	−21.18	COOH	SO_4^{2-}	−1970	−526.6
CH_n	NO_3^-	17.73	3382	COOH	NO_3^-	563.8	2533

Table 10. Fitted MR non-electrolyte-ion and cation-anion interaction parameters b_{ik} (kg/mol) and c_{ik} (kg/mol) for LIFAC. Some parameters are zeros because of lack of experimental data.

i	k	b_{ik}	c_{ik}	i	k	b_{ik}	c_{ik}
H_2O	Na^+	−0.02272	7.5e-6	OH	SO_4^{2-}	0.00302	0.01801
H_2O	NH_4^+	−0.02522	0.00107	COOH	Na^+	0.4478	0.00064
H_2O	H^+	−1.098	0.00154	COOH	NH_4^+	−0.1706	−0.09065
H_2O	Cl^-	0.00073	−0.00683	COOH	H^+	0	0
H_2O	NO_3^-	−0.00276	−0.00272	COOH	Cl^-	0.03495	−0.00022
H_2O	SO_4^{2-}	−0.04705	0.02376	COOH	NO_3^-	0.1014	0.00019
CH_n	Na^+	−0.1891	4.2e-7	COOH	SO_4^{2-}	0.01679	−0.1489
CH_n	NH_4^+	0.01523	−0.01184	Na^+	Cl^-	0.2138	−0.4265
CH_n	H^+	0	0	Na^+	NO_3^-	0.07377	−0.3340
CH_n	Cl^-	−0.3260	2.781	Na^+	SO_4^{2-}	−0.2365	1.938
CH_n	NO_3^-	−0.04466	0.00036	NH_4^+	Cl^-	0.0506	−0.0381
CH_n	SO_4^{2-}	0.04104	−0.09337	NH_4^+	NO_3^-	0.00028	6.7e-5
OH	Na^+	−0.00059	−0.00038	NH_4^+	SO_4^{2-}	0.2566	−1.277
OH	NH_4^+	−0.2618	−1.486	H^+	Cl^-	0.3294	0.00030
OH	H^+	0	0	H^+	NO_3^-	0.2091	−0.3806
OH	Cl^-	5.6e-6	0.2264	H^+	SO_4^{2-}	0.1401	−0.00343
OH	NO_3^-	−0.01297	−0.00134				

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Table 11. Clegg et al. (1992) model parameters B_{ca}, a_{ca}, W_{ca}, U_{ca}, V_{ca}, B^1_{ca}, a^1_{ca} for the IW part of the Ming and Russell model. Parameters for H_2SO_4 and HCl are from Clegg and Brimblecombe (1995) and Clegg et al. (1992) respectively, and the remaining parameters are from Clegg et al. (1998b).

	B_{ca}	a_{ca}	W_{ca}	U_{ca}	V_{ca}	B^1_{ca}	a^1_{ca}
$H^+-SO_4^{2-}$	−46.7149774	9.5	−9.88620169	−5.45640111	−5.98318162	0	0
$H^+-HSO_4^-$	38.2460542	17.0	−11.1152714	−1.03606797	−3.58228743	0	0
$H^+-NO_3^-$	13.5342	17.0	−3.07186	1.96582	−1.41191	0	0
H^+-Cl^-	20.009	13.0	−13.026	−8.806	0	0	0
$NH_4^+-SO_4^{2-}$	−2.858988	13.0	−0.740149	0.940860	−2.587430	0	0
$NH_4^+-NO_3^-$	24.7529	7.0	0.900729	0.379736	−1.42646	−29.9961	13.0
$NH_4^+-Cl^-$	4.65969	15.0	−0.568291	2.07244	−1.25000	0	0
$Na^+-SO_4^{2-}$	34.4660	13.0	−3.72596	−1.95916	−4.86057	0	0
$Na^+-NO_3^-$	26.9994	5.0	0.0526908	0.266644	−2.30288	−21.6050	13.0
Na^+-Cl^-	19.9338	5.0	−5.64608	−3.60925	−2.45982	0	0

Table 12. Fitted UNIFAC organic-ion interaction parameters a_{ij} (K) for the Ming and Russell model. Some parameters are zeros because of lack of experimental data. Water-organic interaction parameters are in the Table 8, and ion-ion and water-ion interaction parameters were set to zero.

i	j	a_{ij}	a_{ji}	i	j	a_{ij}	a_{ji}
CH_n	H^+	0	0	OH	HSO_4^-	0	0
CH_n	NH_4^+	10.76	1235	OH	NO_3^-	−722.1	−229.4
CH_n	Na^+	−715.2	−381.4	OH	Cl^-	2814	738.3
CH_n	SO_4^{2-}	1085	9.192	COOH	H^+	0	0
CH_n	HSO_4^-	0	0	COOH	NH_4^+	−18.89	−195.8
CH_n	NO_3^-	601.5	−256.2	COOH	Na^+	2305	−603.6
CH_n	Cl^-	−790.4	164.2	COOH	SO_4^{2-}	−40.39	−924.3
OH	H^+	0	0	COOH	HSO_4^-	0	0
OH	NH_4^+	−2.412	−385.4	COOH	NO_3^-	−122.8	−5.668
OH	Na^+	−517.1	−480.5	COOH	Cl^-	344.2	5.473
OH	SO_4^{2-}	−15.52	456.8				

Table 13. Fitted UNIFAC non-electrolyte-ion and cation-anion interaction parameters a_{ij} (K) for the Extended UNIFAC. Some parameters are zeros because of lack of experimental data. Water-organic interaction parameters are in the Table 8.

i	j	a_{ij}	a_{ji}	i	j	a_{ij}	a_{ji}
H_2O	Na^+	−174.6	13.00	OH	SO_4^{2-}	0.045	−17.60
H_2O	NH_4^+	−2.326	−0.581	COOH	Na^+	6.684	−0.002
H_2O	H^+	−2157	−355.6	COOH	NH_4^+	16.20	−278.1
H_2O	Cl^-	−7.833	−538.6	COOH	H^+	0	0
H_2O	NO_3^-	−834.5	−192.7	COOH	Cl^-	0.025	5006
H_2O	SO_4^{2-}	−1983	0.001	COOH	NO_3^-	−54.38	−161.1
CH_n	Na^+	2.549	−3.737	COOH	SO_4^{2-}	−2122	0.009
CH_n	NH_4^+	−24.48	0.099	Na^+	Cl^-	−10.59	−640.4
CH_n	H^+	0	0	Na^+	NO_3^-	−1269	36.91
CH_n	Cl^-	5.390	19.54	Na^+	SO_4^{2-}	0.158	−686.1
CH_n	NO_3^-	−1123	0.002	NH_4^+	Cl^-	−894.0	1841
CH_n	SO_4^{2-}	0.537	−0.019	NH_4^+	NO_3^-	−685.4	−489.0
OH	Na^+	6.235	4.452	NH_4^+	SO_4^{2-}	2.168	−438.8
OH	NH_4^+	0.968	−277.7	H^+	Cl^-	−5.281	−3.361
OH	H^+	0	0	H^+	NO_3^-	−0.001	−391.1
OH	Cl^-	0.018	2005	H^+	SO_4^{2-}	143.0	92.51
OH	NO_3^-	−3.648	1.714				

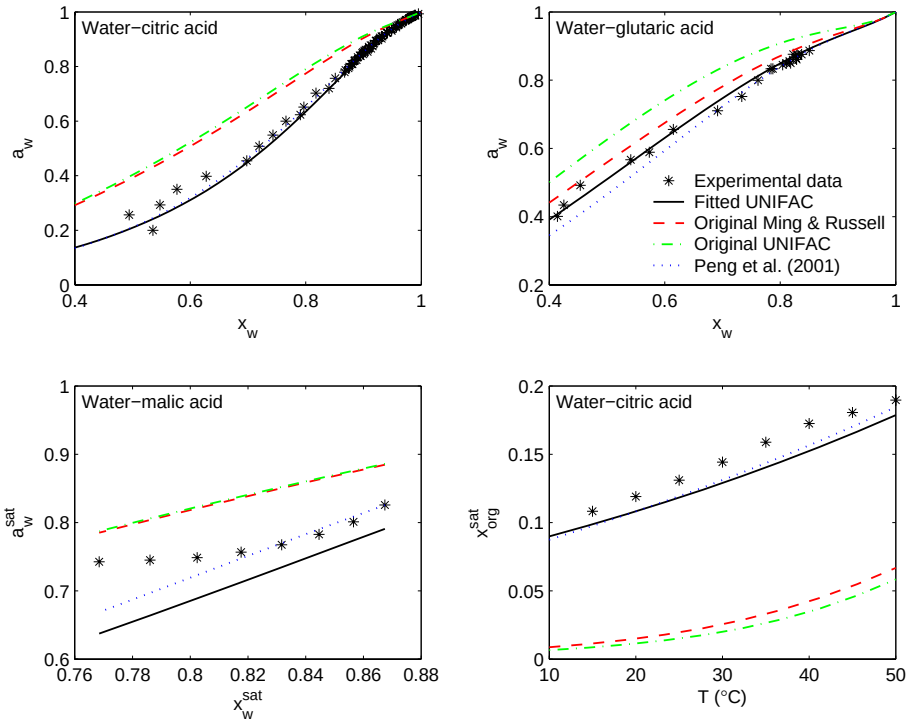

Fig. 1. Experimental and calculated water activities (a_w) in aqueous citric acid and glutaric acid solutions as a function of water mole fraction (x_w) at 298.15 K, experimental and calculated water activities in saturated water-malic acid solution (a_w^{sat}) as a function of water mole fraction (x_w^{sat}) at temperatures from 50°C to 15°C with 5°C intervals, and saturation mole fraction of citric acid (x_{org}^{sat}) as a function of temperature (T). Experimental water activity data for citric acid is from Maffia and Meirelles (2001); Peng et al. (2001); Apelblat et al. (1995a,b); Levien (1955) and data for glutaric acid is from Peng et al. (2001). Solubility and water saturation activity data is from Apelblat et al. (1995a).

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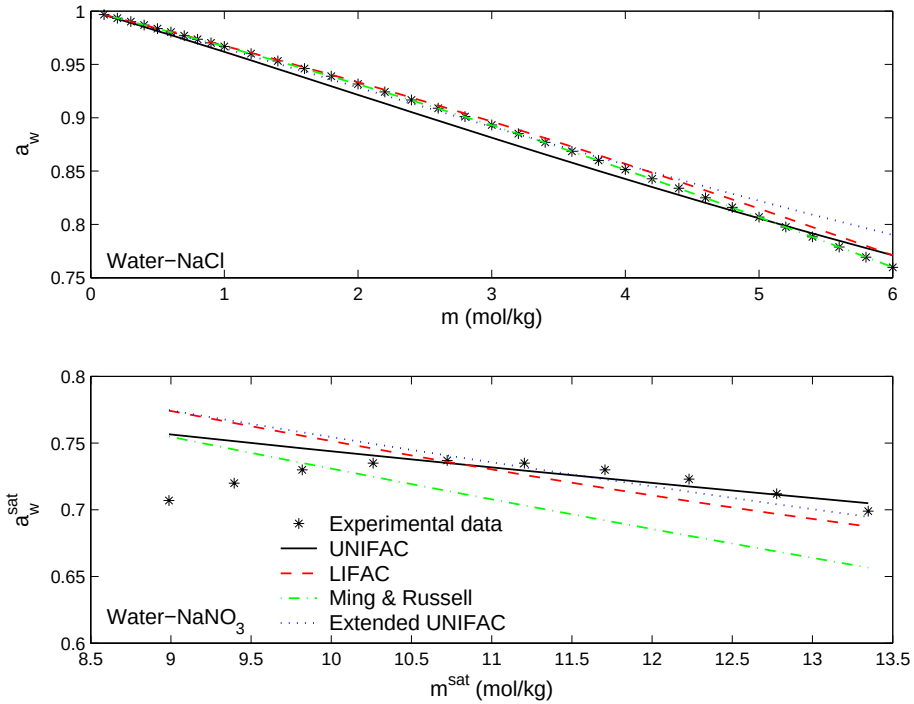

Fig. 2. Experimental and calculated water activities (a_w) in aqueous NaCl solution as a function of molality (m) at 298.15 K, and water activities in saturated water-$NaNO_3$ solution (a_w^{sat}) as a function of saturation molality (m^{sat}) at temperatures from 5°C to 50°C with 5°C intervals. Experimental data for NaCl is from Robinson and Stokes (1959) and for $NaNO_3$ from Apelblat and Korin (1998a).

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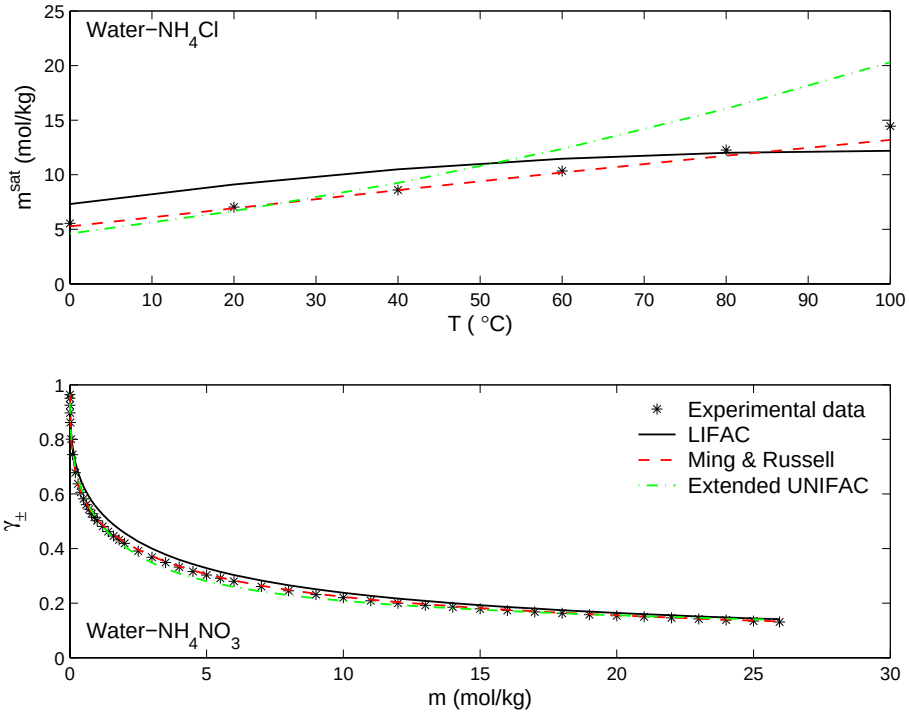

Fig. 3. Experimental and calculated NH_4Cl solubilities (m^{sat}) as a function of temperature (T), and NH_4NO_3 mean activity coefficients ($\gamma_\pm$) as a function of molality (m) at 298.15 K. Experimental solubility data is from ChemDAT – The Merck Chemical Databases (http://chemdat.merck.de/mda/int_en/labtool/index.html) and mean activity coefficient data is from Hamer and Wu (1972).

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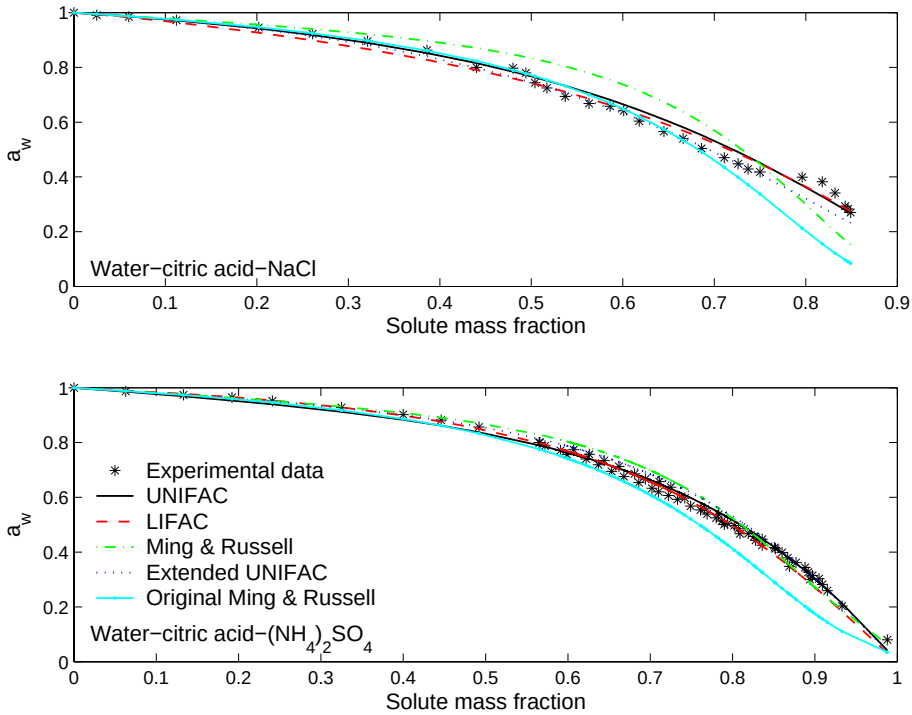

Fig. 4. Experimental and calculated water activities (a_w) in water-organic-electrolyte mixtures as a function of solute mass fraction when organic and salt mole fractions are equal. Experimental data is from Choi and Chan (2002). This data is measured at temperatures between 293.15 and 296.15 K, and activities were calculated using 295 K.

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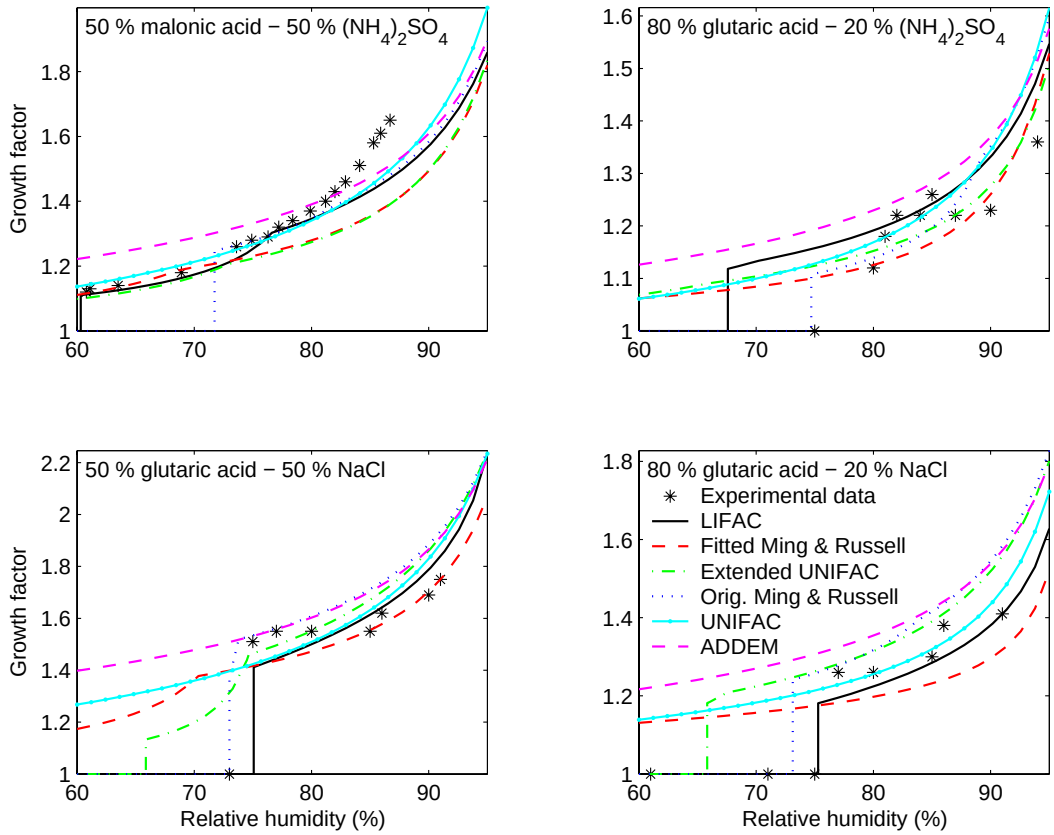

Fig. 5. Experimental and calculated hygroscopic growth factors as a function of relative humidity for four dry 100 nm organic-salt particles. Experimental data for malonic acid-$(NH_4)_2SO_4$ is from Hämeri et al. (2002) and data for other systems is from Cruz and Pandis (2000). The temperature is 298.15 K in Hämeri et al. (2002) data and 297.15 K in Cruz and Pandis (2000) data.

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