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### ► **To cite this version:**

Olivier Bildstein, Christophe Kervévan, Vincent Lagneau, Philippe Delaplace, Anthony Crédoz, et al.. Integrative modeling of caprock integrity in the context of CO<sub>2</sub> storage: evolution of transport and geochemical properties and impact on performance and safety assessment. *Oil & Gas Science and Technology - Revue d'IFP Energies nouvelles*, 2010, 65 (3), pp.485-502. 10.2516/ogst/2010006 . hal-00505305

**HAL Id: hal-00505305**

**<https://minesparis-psl.hal.science/hal-00505305>**

Submitted on 23 Jul 2010

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# Integrative modeling of caprock integrity in the context of CO<sub>2</sub> storage: evolution of transport and geochemical properties and impact on performance and safety assessment

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## 1. ABSTRACT

The objective of the “Géocarbone-Intégrité” project (2005-2008) was to develop a methodology to assess the integrity of the caprock involved in the geological storage of CO<sub>2</sub>. A specific work package of the project (WP5) was dedicated to the integration of (1) the phenomenology describing the evolution of the storage system with a focus on the mechanisms occurring in the caprock and at the interface with the caprock, and (2) the data obtained from the investigation of petrographical, geomechanical, and geochemical properties, before and after reaction with CO<sub>2</sub>-rich solutions, performed in the other work packages (WP1 to WP4). This knowledge was introduced in numerical models and specific safety scenarios were defined in order to assess the performance of the CO<sub>2</sub> storage system.

The results of the modeling show that the injection of CO<sub>2</sub> can potentially have a significant effect on the caprock by changing the porosity due to the dissolution and precipitation of minerals, but that the impact is limited to a zone from several decimeters to several meters of the caprock close to the interface with the reservoir depending on whether the supercritical carbon dioxide (SC-CO<sub>2</sub>) plume enters into the caprock and if fractures are present at this location.

The methodology used in this project can be applied to a pilot site for the injection of CO<sub>2</sub> in the Paris Basin. A key aspect of the safety of such a facility will be to look at the coupling of geochemical alteration and the evolution of geomechanical properties in the short and medium terms (several hundreds of years). The challenge for the future will be to structure and apply the safety assessment methodology with an operational finality, in order to support the robustness of the transition step to CGS projects at the industrial scale.

## 2. RESUME

Le Volet 5 du projet « Géocarbone-Intégrité » visait à intégrer l'ensemble des mécanismes étudiés dans les quatre premiers volets du projet pour une évaluation de performance des couvertures et une étude de sûreté afin de s'assurer de leur préservation et de leur intégrité sur le long terme (de l'ordre du millénaire). L'objectif est d'une part d'aboutir à la construction d'un modèle phénoménologique multi-échelle global, puis à un modèle numérique décrivant le confinement du CO<sub>2</sub> par les couvertures et, d'autre part, de déterminer les performances du confinement en identifiant les processus clefs et les paramètres les plus influents.

Une première partie du programme a consisté en une intégration spatiale de l'ensemble des données phénoménologiques et structurales disponibles à la suite des travaux réalisés dans les différents volets (WP1 à WP4) et à la définition des scénarios types d'évolution du site de stockage (niveaux réservoirs et encaissants). Ce travail a permis de définir les cas tests à

49 prendre en compte et de réaliser les calculs de performance par rapport aux scénarios  
50 d'injection et par rapport aux hétérogénéités majeures identifiées dans les niveaux de  
51 confinement (notamment les fractures).

52 Les résultats montrent que l'injection de CO<sub>2</sub> peut avoir un effet significatif, en  
53 altérant la porosité par dissolution et précipitation de minéraux, mais que l'impact est limité  
54 dans l'espace, de quelques décimètres à quelques mètres de l'interface réservoir-couverture,  
55 selon que la bulle de CO<sub>2</sub> supercritique pénètre ou non dans la couverture et selon la présence  
56 ou l'absence de fractures.

57 La prise en compte des résultats issus de l'analyse de sensibilité et l'analyse des  
58 incertitudes permettra de conduire des calculs de sûreté plus précis. Appliqués au futur site  
59 d'injection, ces calculs permettront d'évaluer la pérennité des propriétés de confinement des  
60 couvertures et de valider la qualité de confinement du site de stockage de CO<sub>2</sub>. Il conviendra  
61 notamment d'évaluer l'impact du couplage entre les phénomènes géochimiques et  
62 géomécaniques sur le court et moyen terme (de l'ordre de la centaine d'année). Le défi pour  
63 l'avenir est de structurer et d'appliquer la méthodologie de l'analyse de sûreté, en mettant en  
64 avant la finalité opérationnelle, de manière à assurer la robustesse de la transition vers les  
65 projets de CGS à l'échelle industrielle.  
66

### 67 3. INTRODUCTION

68

69 The storage of CO<sub>2</sub> in deep saline aquifers and depleted oil and gas reservoirs for  
70 periods of time of ~1,000-10,000 years is considered in order to mitigate its release in the  
71 atmosphere and avoid the consequences of the additional greenhouse effect on climate change  
72 (IPCC, 2005). The feasibility of such an industrial process and the safety on the long term has  
73 to be demonstrated and relies mainly on the confinement properties of the caprock. In general,  
74 the knowledge of the structure, the properties and the reactivity of the caprock is poor because,  
75 usually, the reservoir is the main object of interest for oil and gas production.

76 The objective of the "Geocarbone-Intégrité" project (2005-2008) was therefore to  
77 develop a methodology and to design a tool to assess the integrity of the caprock involved in  
78 the geological storage of CO<sub>2</sub>. A specific work package of the project (WP5) was dedicated to  
79 the integration of (1) the phenomenology describing the evolution of the storage system with  
80 a focus on the mechanisms occurring in the caprock and at the interface with the caprock, and  
81 (2) the data obtained from the investigation of petrographical, geomechanical, and  
82 geochemical properties, before and after reaction with CO<sub>2</sub>-rich solutions, performed in the  
83 other work packages (WP1 to WP4) (see Fleury et al., *this issue*, for a detailed description of  
84 the project). The ultimate goal is to construct a conceptual and numerical model at the site  
85 scale to predict the evolution of the storage on the long term and to ensure the persistence of  
86 the caprock integrity. This model is developed in the perspective of the assessment of the  
87 performance and safety of the future injection pilot site in the Paris Basin planned to be  
88 commissioned in 2010.

89 A review of the existing literature on CO<sub>2</sub> storage modeling reveals that most of the  
90 effort made by the scientific community are devoted to the study of injectivity properties and  
91 mineral trapping capability in reservoirs (see review by Gaus et al., 2008). The studies on  
92 caprock integrity are still relatively rare, and only recently some insights on caprock  
93 mineralogical alteration patterns induced by CO<sub>2</sub> migration have been gained by means of  
94 reactive transport modeling techniques (Johnson et al., 2004, 2005; Gauss et al., 2005; Xu et  
95 al., 2005; Gherardi et al. 2007). This study is focused on the numerical prediction of the long  
96 term variations of the mineralogical and hydraulic properties of the caprock in the French  
97 pilot site for CO<sub>2</sub> geological storage, in the Paris Basin.

## 100 4. PHENOMENOLOGY OF THE STORAGE

101

### 102 4.1. Physicochemical processes at the interface with the caprock

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104 The phenomenology of the storage is described in detail in a special report of the  
105 Intergovernmental Panel on Climate Change dedicated to the capture and storage of CO<sub>2</sub>  
106 (IPCC, 2005). The injected CO<sub>2</sub> is usually at supercritical conditions (SC-CO<sub>2</sub>) in the typical  
107 reservoir pressure and temperature conditions (63°C and 145 bar in the case of the Saint  
108 Martin-de-Bossenay field - Paris basin, France). Under these conditions, the fluid properties  
109 of SC-CO<sub>2</sub> are similar to both a liquid phase (density around 0.6) and a gas phase (low  
110 viscosity; e.g. Mathias et al. 2009). Also, CO<sub>2</sub> is very soluble in water: about 1 mol/l (e.g.,  
111 Duan and Sun, 2003). Its migration in porous media (reservoirs and caprocks) containing  
112 water involves capillary effects.

113 Since SC-CO<sub>2</sub> is less dense than water, it will rise in the reservoir. A fraction of this  
114 CO<sub>2</sub> will be trapped in the porosity (capillary trapping) and the rest will reach the structural  
115 trap (or stratigraphical trap) constituted by the caprock, which is expected to prevent the CO<sub>2</sub>  
116 from rising any further and eventually reaching the atmosphere. This is due to the properties  
117 of the caprock which is usually a clay-rich material, saturated with water and characterized by  
118 a very low permeability and a high gas entry pressure. The caprock will therefore be in  
119 physical contact with the SC-CO<sub>2</sub> plume during most of the storage lifetime.

120 If the overpressure of SC-CO<sub>2</sub> is lower than the capillary entry pressure, the SC-CO<sub>2</sub>  
121 will remain confined in the reservoir. However, dissolved CO<sub>2</sub> will still be able to diffuse into  
122 the caprock. This is a slow transport process but the dissolution of CO<sub>2</sub> can strongly affect the  
123 composition of the formation water, in particular, by lowering the local pH. This change can  
124 potentially damage the caprock by destabilizing the chemical equilibrium with the primary  
125 mineral phases and triggering the dissolution of some of them and the precipitation of  
126 secondary phases.

127 If the overpressure at the interface between the reservoir and the caprock overcomes the  
128 entry pressure, the SC-CO<sub>2</sub> will penetrate into the caprock, due to the pressure gradient and  
129 the buoyancy forces, and will displace the caprock water. If the pressure further builds up, the  
130 plume can potentially force its way, in mechanical terms, into the caprock through dilatancy  
131 driven flow or induced fracturing.

132

133 Importantly, any heterogeneity in the caprock, such as small cracks or fractures, will  
134 facilitate the migration of SC-CO<sub>2</sub> into the caprock. The behavior of these preferential  
135 pathways and the reactivity with the CO<sub>2</sub>-rich fluids is critical for the understanding of the  
136 evolution of the confinement properties of the caprock. To this regard, the dissolution of the  
137 mineral phase constituting the cement of the rock (e.g. carbonates) can potentially open or  
138 close the porosity and affect the permeability of the matrix as well as of the cracks and  
139 fractures. The same effect can be obtained by altering potentially expansive primary clay  
140 minerals and forming secondary non-swelling ones (e.g. through the illitization process;  
141 Crédoz *et al.*, 2009).

142 A last potentially important migration pathway for CO<sub>2</sub>, which was not investigated in  
143 this work, is the possibility of having defective abandoned wells present in the zone  
144 influenced by the injection.

145

146

## 147 4.2. Scenarios for the evolution of the storage

148 The following scenarios have been considered to predict the fate of SC-CO<sub>2</sub> at the  
149 interface between reservoir and caprock (Figure 1):

- 150 1. The SC-CO<sub>2</sub> overpressure at the top of the reservoir is **lower than the**  
151 **capillary entry pressure in the caprock**. As a consequence, the SC-CO<sub>2</sub>  
152 cannot penetrate into the caprock but dissolved CO<sub>2</sub> and acidified formation  
153 water can penetrate into the caprock by diffusion, triggering geochemical  
154 alteration. This is the *reference case scenario* for the safety assessment  
155 (section 5.2.1, *case 1a*, *case 1b*, and *case 1c*).
- 156 2. The SC-CO<sub>2</sub> does not directly enter into the rock matrix but penetrates the  
157 caprock through a network of connected fractures. This is the *"fracture*  
158 *network" scenario* which is considered as probable and constitutes a first  
159 altered scenario for the safety assessment. Only a dissolved CO<sub>2</sub> is considered  
160 in this fracture scenario (section 5.2.2, *case 2*).
- 161 3. The SC-CO<sub>2</sub> overpressure at the reservoir top is **higher than the capillary**  
162 **entry pressure in the caprock**. In this case, the SC-CO<sub>2</sub> enters into the  
163 caprock by forced drainage. This is the *multiphase scenario* where the SC-  
164 CO<sub>2</sub> migration is controlled by the effective caprock permeability. This  
165 scenario is considered as highly probable in the injection phase (due to  
166 significant overpressure) and constitutes a second altered scenario for the  
167 safety assessment (section 5.3, *case 3*).

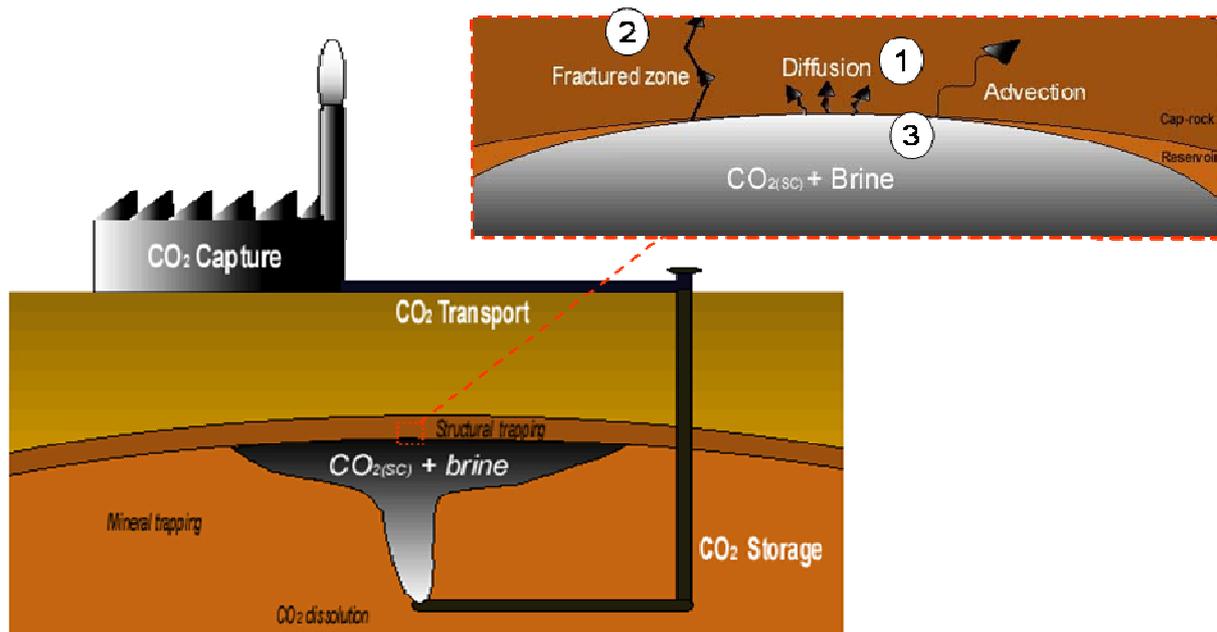
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169 The chosen scenarios mainly focus on the geochemical effects induced by the  
170 advancement of both a CO<sub>2</sub>-rich aqueous phase and a free SC-CO<sub>2</sub> gas plume through the  
171 caprock. A fractured caprock scenario is also simulated but the focus is concentrated on  
172 geochemical reactions with the filling material (calcite). In this work, worst case scenarios  
173 based on "hydrodynamic" leaking of CO<sub>2</sub>, e.g. through open fractures, are not investigated  
174 and therefore the results presented here does not constitute a full safety assessment exercise.

175

176 The effect of gas-rock interactions has not been considered in our calculations because  
177 most of the reactivity is expected to occur at the interface between rock and aqueous phase.  
178 This is supported by the fact that high residual water contents (liquid saturation,  $S_L$ , always  
179 greater than about 0.6) have been calculated even for the case of most significant penetration  
180 of SC-CO<sub>2</sub> into the caprock.

181



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184 **Figure 1. Scenarios for the evolution of the reservoir/caprock system during CO<sub>2</sub> injection.**

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## 186 **5. ASSESSMENT OF CAPROCK INTEGRITY PERSISTENCE WITH TIME**

187

188 The impact of geochemical alteration of the caprock, in terms of confinement properties,  
189 on transport properties such as permeability depends primarily on the initial value of these  
190 parameters and on the net volume balance of mineral reactions, i.e. dissolution and  
191 precipitation (Bemer and Lombard, 2009).

192 The geochemical reactivity of the caprock and its constitutive minerals has been  
193 investigated in WP4 of the “Géocarbonate-Intégrité” project which aimed at determining the  
194 reaction pathways for SC-CO<sub>2</sub>/water/rock interactions and also kinetic parameters for  
195 carbonate and clay mineral transformations (Kohler and Parra, 2007; Crédoz *et al.*, 2009;  
196 Hubert, 2009).

197 These parameters have to be integrated into large scale modeling in order to calculate  
198 the evolution of the storage system as a result of the CO<sub>2</sub> perturbation. A major challenge of  
199 this type of modeling is to extrapolate the behavior of the system from the lab scale to the  
200 field scale, including:

- 201 – time: 1 year for experiments vs. 10,000 years for geological storage,
- 202 – solid/solution ratio: around 50 g/L in experiments vs. about 50 kg/L in depth,
- 203 – texture: fine crushed rock (0.1 mm particles) vs. bulk rock,
- 204 – temperature: 80-150°C in the experiments vs. 65°C in Saint Martin-de-Bossenay  
205 (Paris Basin, France).

206

207 In the laboratory, the conditions are chosen so that the reactivity of the samples is  
208 enhanced in order to limit the duration of the experiments. For the large scale calculations, it  
209 is crucial to set these parameters to realistic values but the effect of the texture is difficult to  
210 determine except through values of the reactive surface area. Petrophysical and geochemical  
211 parameters are given for the different scenarios considered in Table 1 which are used for 1D  
212 and 2D calculations with a simplified caprock geometry (15% porosity, 1 to 10 meters thick).

213

## 5.1. Modeling approach

At the beginning of the project, the collaborative work between different modeling teams was organized much in the same way as the IPCC modeling of climate (IPCC, 2007), i.e. defining a set of common guidelines and parameters for the modeling (Table 1) but also giving some degree of liberty concerning the way to carry out the calculations, in particular:

- the choice of the numerical tool,
- the capillary and permeability properties,
- the list of secondary minerals allowed to precipitate, and to some extent of primary minerals as well,
- the values for the mineral kinetic constants for precipitation and dissolution and the reactive surface area,
- the feedback between mineral dissolution/precipitation and the transport parameters (diffusion coefficient, permeability, capillary curve, ...)

In this way, the modeling exercise should not be considered as a benchmarking of different modeling tools but rather as an investigation of the dominant processes and the most influential parameters giving an envelope of behaviors for the storage system. Only one representative set of results is shown for a specific scenario if all the modeling teams involved in the calculations reached the same conclusions. If the conclusion is significantly different, a comparison and analysis of the results is presented.

### *Guidelines for the modeling scenarios*

A first series of calculations in scenarios where the porous media are saturated (Table 1: cases 1a to 1c and case 2, respectively corresponding to scenarios 1 and 2) were performed with reactive transport tools available in the different modeling teams: Crunch (Steeffel, 2001), Hytec (van der Lee et al., 2003), PhreeqC (Parkhurst and Appelo, 1999), PHAST (Parkhurst et al. 2004). The thermodynamic database used for the calculations is derived from EQ3/6 code (Wolery, 1992), and the kinetic data for the dissolution (and to a lesser degree for the precipitation) of mineral phases were taken from the review by Palandri and Kharaka (2004).

A second series of calculations using the same scenarios involved multiphase flow and reactive transport in porous media (Table 1: case 3 corresponding to scenario 3) and were performed with TOUGHREACT (Xu and Pruess, 2001) and COORES<sup>TM</sup> (e.g. Le Gallo et al. 2007).

Reference case		Sensitivity analysis
Duration = 10000 years - Temperature = 80°C		
Caprock initial composition	based on Charmotte/Saint Martin-de-Bossenay (Paris basin)	
Water initial composition	in equilibrium with caprock mineralogy (pH = 6.5)	
Boundary condition (constant concentration)	(1a) acidified water starting from Dogger formation $\text{CO}_2(\text{aq}) = 1.1$ molal in equilibrium with $\text{pCO}_2 = 150$ bar (pH = 4.7) (1b) initial composition acidified with $\text{pCO}_2 = 150$ bar and buffered with carbonates (pH = 4.6) (1c) initial composition acidified with $\text{pCO}_2 = 150$ bar (pH = 3.4)	(1a) water from Dogger formation (pH = 6.2)  (1b) initial composition acidified with $\text{pCO}_2 = 150$ bar (pH = 3.4)
<b>1) 1D Diffusive/convection case (1a, 1b, 1c)</b>		

Porosity	15%	(1a) 5%
Effective diffusion coefficient	$10^{-11} \text{ m}^2/\text{s}$	(1a) $10^{-10} \text{ m}^2/\text{s}$
Permeability Flow rate	(1c) $K = 1.6 \cdot 10^{-18} \text{ m}^2$ 10x diffusive flux	
<b>2) 2D system with discrete fracture (case 2)</b>		
Mineralogy and water composition	based on case 1c	
Fracture filled with calcite	porosity 40%	
Fracture permeability	10,000 x higher than in reservoir (case 1c)	
<b>3) 1D multiphase case (case 3)</b>		
Mineralogy and water composition	based on single-phase case 1a	
Boundary condition	constant pressure	
Relative permeability	Van Genuchten model (see Eq. 1 and 2) $K = 10^{-18} \text{ m}^2$	
Capillary pressure	Van Genuchten model (see Eq. 3)	
Effective diffusion coefficient	$10^{-11} \text{ m}^2/\text{s}$	

249 **Table 1. Modeling parameters for the reference case simulation and sensitivity analyses.**

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Note that for the boundary conditions, the chemical composition of reservoir pore waters remain fixed during the simulation. For the cases with acidified waters, it means that the pH in the reservoir is controlled by the CO<sub>2</sub>-plume during 10,000 years, even though control should be taken over by the reservoir water composition again at some point after the end of the CO<sub>2</sub> injection. It is however considered here as a conservative assumption for the performance and safety of the storage.

## 258 **5.2. Saturated caprock: geochemical interactions with dissolved CO<sub>2</sub>**

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In these scenarios, the SC-CO<sub>2</sub> plume is trapped in the reservoir and the acidic perturbation migrates by diffusion of dissolved species only.

### 262 **5.2.1. Homogeneous caprock, diffusive/advective case**

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In these calculations, the diffusion coefficient is set initially to  $10^{-11} \text{ m}^2/\text{s}$ , which is the mean value measured in the argillites from Bure (Talandier et al., 2006) and considered as analogues to the clay series in the caprock investigated in the framework of this project. More details about diffusion coefficients can be found in Fleury et al. (2009) and Berne et al. (2009). Some sensitivity calculations are also shown which investigate the influence of this parameter.

#### 270 ***Case 1a: pure diffusion, specific water compositions for the reservoir and caprock***

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In this case, the caprock is initially homogeneous in composition (mineralogy and pore water) and in transport properties (no cracks or fractures or other preferential pathways). Preliminary batch calculations were conducted in order to determine possible secondary mineral phases that potentially precipitate in this context. Some simplifications were made in the calculations in order to avoid dealing with complex solid-solutions: clay minerals are represented only by pure end-members (it concerns especially the interstratified illite-smectite minerals), mixed carbonates such as ankerite are considered with a constant composition (Table 3). In this series of simulations, the acidified water ( $p\text{CO}_2 = 150 \text{ bar}$ ) used as the boundary condition at the contact with the caprock is equilibrated with the mineral assemblage of the reservoir. Under this hypothesis, the water can be considered as less

281 aggressive with respect to the caprock minerals than in the configuration where the water is  
 282 only acidified by CO<sub>2</sub> (*case Ib*). The water compositions considered here, based upon data  
 283 from Azaroual et al. (1997), are detailed in Table 2.

284 A 1D geometry was considered within the caprock assuming an initial equilibrium  
 285 between the water (Table 2, 3<sup>rd</sup> column) and the mineral phases constituting the caprock  
 286 (Table 3). The base of the modeled domain was supposed to be permanently in contact with  
 287 the acidified reservoir water (Table 2, 2<sup>nd</sup> column) so that dissolved CO<sub>2</sub> is transported by  
 288 molecular diffusion within the caprock.

289 We performed several simulations considering various typical initial porosities (15 and  
 290 5%) and diffusion coefficient values (10<sup>-11</sup> and 10<sup>-10</sup> m<sup>2</sup>.s<sup>-1</sup>). The codes PHREEQC and  
 291 PHAST were used with various meshes (about 100 grid cells, with both uniform and variable  
 292 grid spacing) and time-stepping in order to increase the robustness of the calculations  
 293 presented hereafter (Figure 2 and Figure 3).

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Reference Dogger reservoir water (80°C)*		Acidified reservoir water (80°C, pCO <sub>2</sub> = 150 bar)*		Initial caprock water (80°C)	
pH	6.24	pH	4.75	pH	6.54
Species	Molality	Species	Molality	Species	Molality
Al	5.622 10 <sup>-8</sup>	Al	1.251 10 <sup>-7</sup>	Al	1.531 10 <sup>-7</sup>
C	4.895 10 <sup>-3</sup>	C	1.141	C	2.180 10 <sup>-3</sup>
Ca	1.612 10 <sup>-2</sup>	Ca	3.204 10 <sup>-2</sup>	Ca	1.528 10 <sup>-2</sup>
Cl	3.014 10 <sup>-1</sup>	Cl	3.015 10 <sup>-1</sup>	Cl	2.601 10 <sup>-1</sup>
Fe	2.137 10 <sup>-7</sup>	Fe	1.751 10 <sup>-6</sup>	Fe	1.534 10 <sup>-5</sup>
K	2.374 10 <sup>-3</sup>	K	2.375 10 <sup>-3</sup>	K	1.190 10 <sup>-2</sup>
Mg	1.282 10 <sup>-2</sup>	Mg	2.424 10 <sup>-2</sup>	Mg	8.937 10 <sup>-4</sup>
Na	2.594 10 <sup>-1</sup>	Na	2.595 10 <sup>-1</sup>	Na	2.543 10 <sup>-1</sup>
S	7.642 10 <sup>-3</sup>	S	7.649 10 <sup>-3</sup>	S	1.841 10 <sup>-2</sup>
Si	8.994 10 <sup>-4</sup>	Si	8.833 10 <sup>-4</sup>	Si	5.371 10 <sup>-4</sup>

296 (\*) equilibrated with the assumed reservoir mineralogy (calcite, dolomite-dis, chalcedony, illite, pyrite)

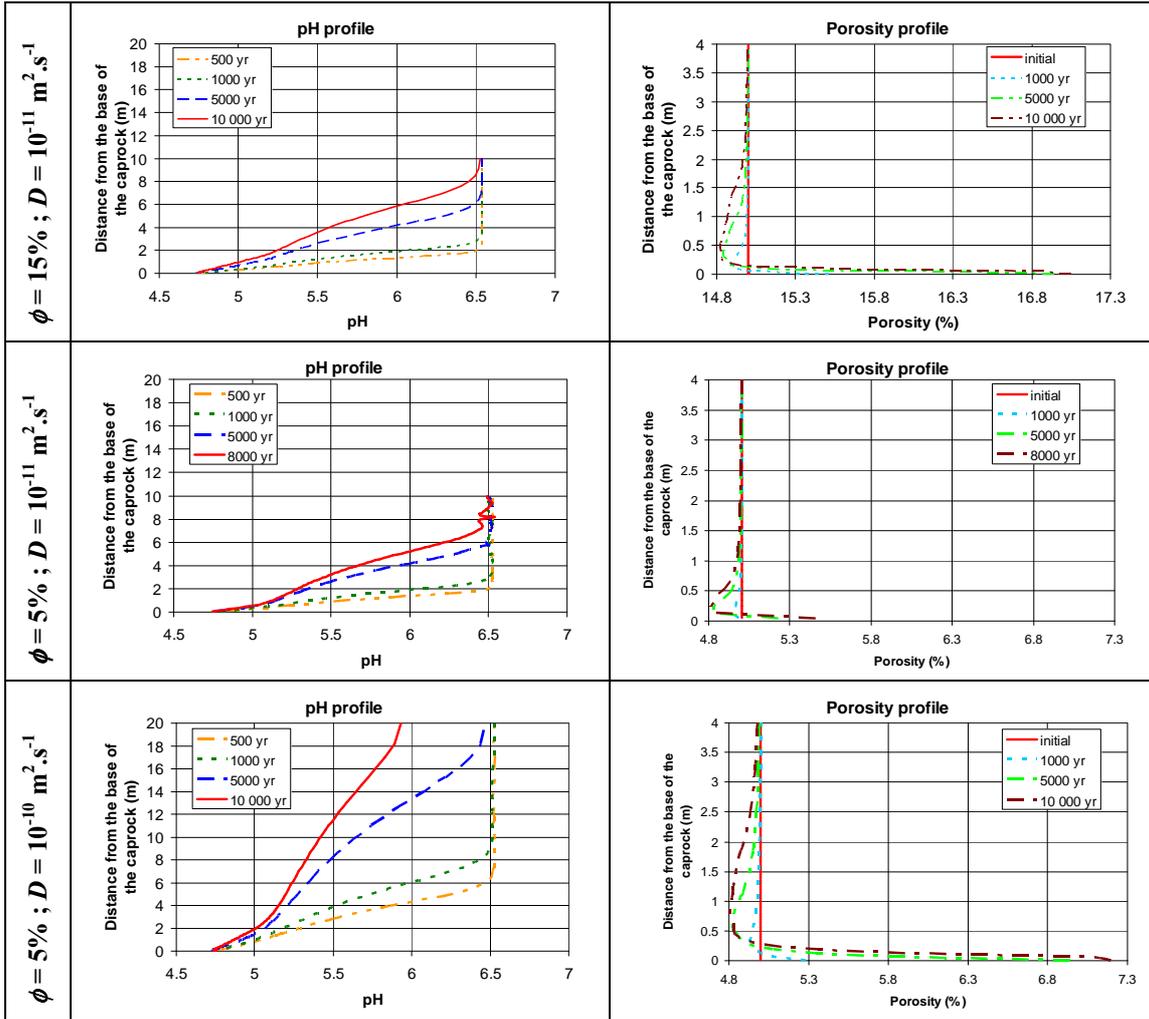
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298 **Table 2. Water compositions considered in the simulations.**

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Mineral phase	Initial weight %	log(K) at 80°C	reaction	SI(1)	SI(2)	SI(3)
Calcite	50	1.05	Calcite + H <sup>+</sup> = Ca <sup>++</sup> + HCO <sub>3</sub> <sup>-</sup>	0	0	0
Ankerite	5	12.14	Ankerite + 4H <sup>+</sup> = Ca <sup>++</sup> + 0.3Mg <sup>++</sup> + 0.7Fe <sup>++</sup> + 2H <sub>2</sub> O + 2CO <sub>2(aq)</sub>	-0.99	-0.99	0
Montmorillonite-Na	25	-0.65	Mont-Na + 6H <sup>+</sup> = 0.33 Mg <sup>++</sup> + 0.33 Na <sup>+</sup> + 1.67Al <sup>+++</sup> + 4H <sub>2</sub> O + 4SiO <sub>2(aq)</sub>	0.77	0.52	0
Kaolinite	3	2.38	Kaolinite + 6H <sup>+</sup> = 2Al <sup>+++</sup> + 2SiO <sub>2(aq)</sub> + 5H <sub>2</sub> O	0.13	1.49	0
Illite	2	3.80	Illite + 8H <sup>+</sup> = 0.25Mg <sup>++</sup> + 0.6K <sup>+</sup> + 2.3Al <sup>+++</sup> + 3.5SiO <sub>2(aq)</sub> + 5H <sub>2</sub> O	0	0	0
Quartz	10	-3.24	Quartz = SiO <sub>2(aq)</sub>	0.23	0.23	0
Anhydrite	3	-5.05	Anhydrite = Ca <sup>++</sup> + SO <sub>4</sub> <sup>--</sup>	-0.47	-0.34	0
Pyrite	2	-21.91	Pyrite + H <sub>2</sub> O = 0.25H <sup>+</sup> + 0.25SO <sub>4</sub> <sup>-</sup> + Fe <sup>++</sup> + 1.75HS <sup>-</sup>	0	0	0
<i>Goethite</i>	-	-1.13	Goethite + 3H <sup>+</sup> = Fe <sup>+++</sup> + 2H <sub>2</sub> O	-2.65	-5.20	6e-4
<i>Chalcedony</i>	-	-3.02	Chalcedony = SiO <sub>2(aq)</sub>	0	0	-0.23
<i>Disordered-Dolomite</i>	-	1.92	Dolom-dis + 2H <sup>+</sup> = Ca <sup>++</sup> + Mg <sup>++</sup> + 2HCO <sub>3</sub> <sup>-</sup>	0	0	-1.24
<i>Siderite</i>	-	-1.17	Siderite + H <sup>+</sup> = Fe <sup>++</sup> + HCO <sub>3</sub> <sup>-</sup>	-2.78	-2.78	-0.84

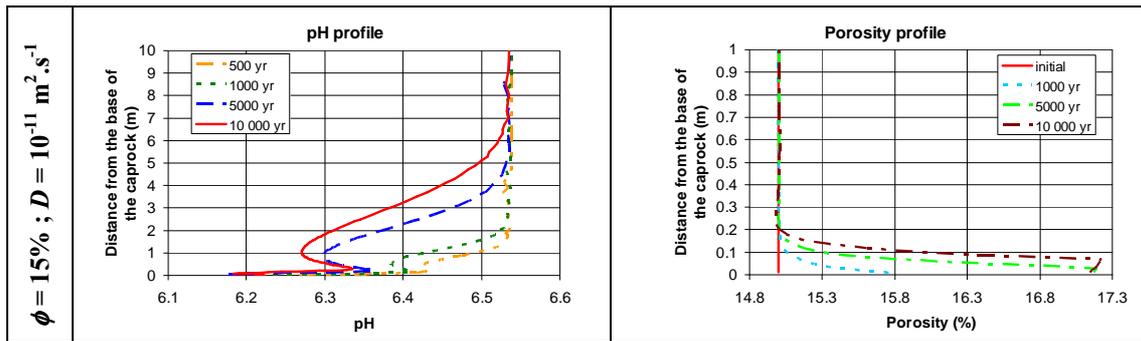
300 **Table 3. Initial mineral composition of the caprock (inspired from the Charmotte field) and**  
 301 **secondary phases (*in italics*) taken into account in the model. Equilibrium constants at 80°C and**  
 302 **reactions are given for each mineral. SI(1), SI(2), and SI(3) refer to the initial saturation indexes**  
 303 **of the mineral phases in the Dogger, acidified, and caprock waters respectively.**



**Figure 2. Diffusion of the acidified reservoir water: pH and porosity profiles in the caprock calculated for several porosity (15 and 5%) and diffusion coefficient ( $10^{-11}$  and  $10^{-10} \text{ m}^2 \cdot \text{s}^{-1}$ ) values. (Note zoom over 4 m in porosity profiles)**

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In all cases, similar trends and orders of magnitude for pH and porosity profiles along the first few meters of the caprock are observed after a simulation period up to ten thousands years (Figure 2). As expected, the acidified water penetrates the caprock on a distance increasing with the diffusion coefficient value (see pH profiles). However, the impact on the minerals remains comparable (mainly: dissolution of illite and anhydrite, precipitation of calcite, montmorillonite-Na and kaolinite), showing no significant sensitivity towards initial porosity and diffusion coefficient values, as illustrated by the porosity profiles which integrate the volumetric variations of all the primary and secondary mineral phases considered in the chemical reactions (see Table 3). In all cases, both the amplitude and the extension of the variations observed along the modeled profile remain relatively limited: a first increase in porosity (up to +2.2%) within the first decimeters of the caprock domain is observed, followed by a slight decrease of the porosity value (around -0.2%) in the next 1 to 3 meters.



322 **Figure 3. Diffusion of the non-acidified reservoir water: pH and porosity profiles in the caprock**  
 323 **calculated for a porosity of 15% and a diffusion coefficient of  $10^{-11} \text{ m}^2 \cdot \text{s}^{-1}$ .**

324  
 325 Note that the water compositions described in Table 2 also show a significant contrast  
 326 between the initial caprock water and the original reservoir water. Formation water indeed  
 327 often show differences which are not completely balanced by diffusion at the interface.  
 328 Interdiffusion processes can potentially induce some transformations in the mineral  
 329 composition at this location. In order to quantify this effect and discriminate the role of  
 330 dissolved  $\text{CO}_2$ , we performed another simulation for the first parameter set ( $\phi = 15\%$ ;  
 331  $D = 10^{-11} \text{ m}^2 \cdot \text{s}^{-1}$ ) now using columns 1 and 3 in Table 2 as boundary and initial water  
 332 compositions, respectively. The results obtained are presented in Figure 3. As expected, pH  
 333 variations along the profile are much smaller in amplitude than in the “acidified water” case  
 334 (ranging here between 6.2 and 6.5). The behavior of the system, in terms of mineral reactivity,  
 335 is comparable to the previous simulations with acidified water (*i.e.*, dissolution of illite and  
 336 anhydrite, precipitation of calcite, montmorillonite-Na and kaolinite) but remains limited to  
 337 the very first decimeters of the caprock. However, the impact on porosity is significant as  
 338 shown by the porosity profiles in Figure 3. When compared to the profiles in Figure 2, it can  
 339 be noticed that the amplitude of the increase in porosity (varying from 15 to 17%) is quite  
 340 similar. In this case, however, no decrease of the porosity is observed in the first meter of the  
 341 caprock.

342  
 343 The results obtained in this last simulation give us some new insight into the specific role of  
 344 the initial water composition and the pH perturbation due to  $\text{CO}_2$ . The distinct impact of the  
 345 initial water was also observed in experiments where the caprock from the Paris Basin was  
 346 reacted with typical reservoir water from the Dogger formation (Crédoz et al., 2009).

347  
 348 ***Case 1b: pure diffusion, same initial water composition in reservoir and caprock***

349 The mineral composition corresponds to that of the transition zone between the  
 350 reservoir and the caprock in the Charmotte area (Paris Basin, France) (Table 4). The initial  
 351 water composition is in equilibrium with the mineral assemblage of the caprock. Two water  
 352 compositions are considered at the boundary (*i.e.* at the interface between the caprock and the  
 353 reservoir). In both case, the water is similar to the previous one, except for the acidification of  
 354 the water due to the dissolution of  $\text{CO}_2$  in the reservoir (Table 5). For the first one, we  
 355 consider the acidification effect but also some short term buffering capacity of minerals  
 356 (essentially carbonates and sulfates) which reacted with the  $\text{CO}_2$  plume and the solution  
 357 during the migration of the plume in the reservoir. This case resembles the previous *case 1a*  
 358 ( $\text{pH} = 4.63$ ). For the second one, we consider that the acidification is maximized ( $\text{pH} = 3.36$ ,  
 359 in equilibrium with  $\text{CO}_2$ -SC only) in order to obtain the strongest pH perturbation possible in  
 360 this system, *i.e.* testing extremely adverse conditions for safety assessment purposes. This  
 361 scenario corresponds to a hypothetical case in which the minerals from the reservoir do not

362 buffer the pH (i.e., the residence time of the interstitial water is much shorter than the reaction  
 363 kinetics characteristic time).  
 364  
 365

	Volume % (fraction of total rock volume)	Weight % (of solid)
<i>Porosity</i>	15	-
<b>Clay fraction</b>		
Illite	11	13
Montmor-Ca	2	3
Kaolinite	9	10
<b>Silt fraction</b>		
Quartz	9	10
<b>Carbonate fraction</b>		
Calcite	41	45
Dolomite	2	3
Siderite	1	2
<b>Accessory minerals</b>		
Pyrite	2	4
Anhydrite/gypsum	4	5
Anatase, other ...*	4	5

\* considered as inert minerals

366  
 367 **Table 4. Simplified composition used for calculations with the transition zone between reservoir**  
 368 **and caprock at Charmotte (porosity = 15%).**

369  
 370

Charmotte transition zone water composition (80°C)*		Acidified water (80°C, pCO <sub>2</sub> = 150 bar, fCO <sub>2</sub> = 83 bar)		Acidified water equilibrated with carbonates* (80°C)	
pH	6.20	pH	3.36	pH	4.63
Species	Molality	Species	Molality	Species	Molality
Al	2.18 10 <sup>-8</sup>	Al	2.18 10 <sup>-8</sup>	Al	1.21 10 <sup>-8</sup>
C	2.00 10 <sup>-3</sup>	C	1.00	C	1.02
Ca	4.22 10 <sup>-2</sup>	Ca	4.22 10 <sup>-2</sup>	Ca	5.50 10 <sup>-2</sup>
Cl	1.92 10 <sup>-1</sup>	Cl	1.92 10 <sup>-1</sup>	Cl	1.92 10 <sup>-1</sup>
Fe	1.22 10 <sup>-5</sup>	Fe	1.22 10 <sup>-5</sup>	Fe	9.30 10 <sup>-6</sup>
K	1.85 10 <sup>-2</sup>	K	1.85 10 <sup>-2</sup>	K	1.85 10 <sup>-2</sup>
Mg	1.12 10 <sup>-2</sup>	Mg	1.12 10 <sup>-2</sup>	Mg	1.52 10 <sup>-2</sup>
Na	8.00 10 <sup>-2</sup>	Na	8.00 10 <sup>-2</sup>	Na	8.00 10 <sup>-2</sup>
S	6.16 10 <sup>-3</sup>	S	6.16 10 <sup>-3</sup>	S	5.86 10 <sup>-3</sup>
Si	6.55 10 <sup>-4</sup>	Si	6.55 10 <sup>-4</sup>	Si	2.48 10 <sup>-4</sup>

\* equilibrated with fast reacting minerals assumed to be present in reservoir (calcite, dolomite, anhydrite)

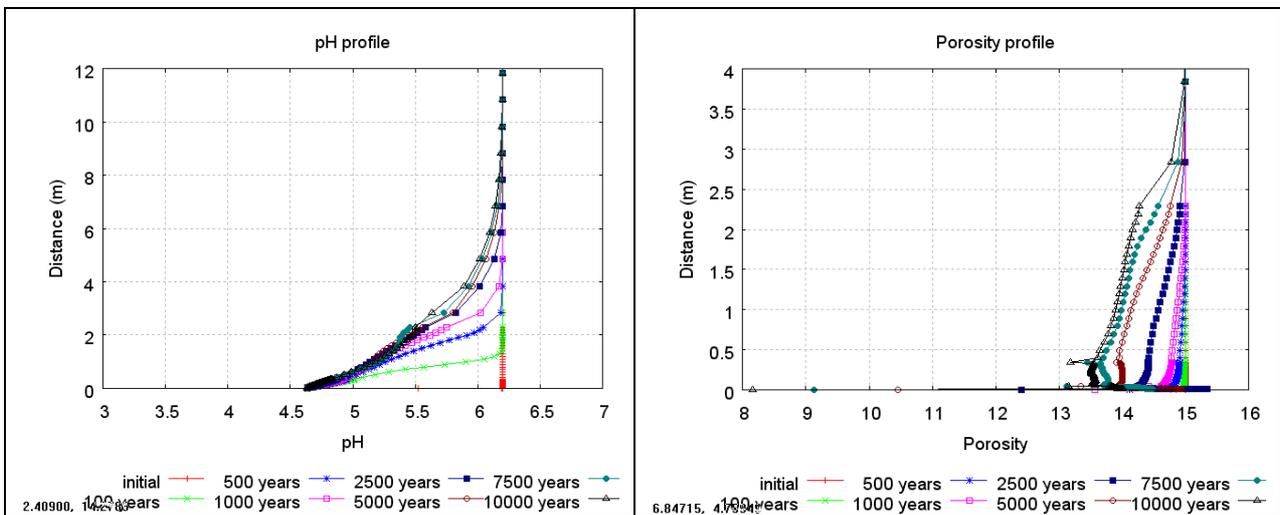
371  
 372

373 **Table 5. Water compositions considered in the simulations.**

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 375  
 376  
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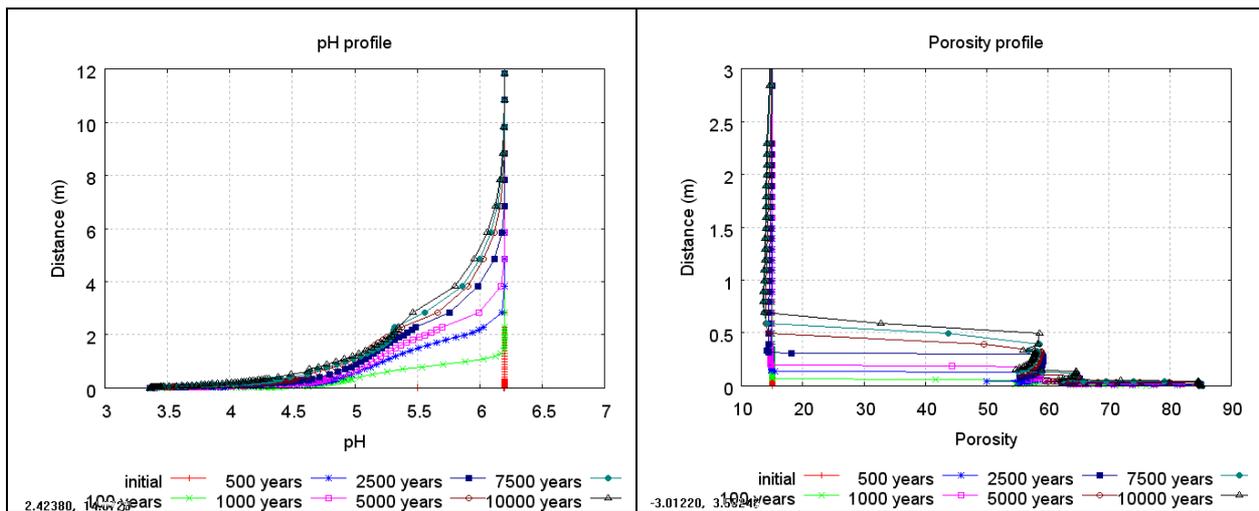
The calculations concerning the scenario with the pH 4.6 water entering the caprock shows that the alteration of the caprock is significant with a similar front of pH perturbation compared to *case Ia*, but the porosity change is very limited (Figure 4). The precipitation of anhydrite (+16%), kaolinite (+12%) and quartz (+2%) is responsible for the porosity decrease

379 observed directly at the reservoir-caprock interface (down to 8% porosity in the first  
 380 centimeter). Illite (-11%), calcite (-6%), and montmorillonite (-2%) dissolution is also  
 381 observed at this location. Calcite dissolution and anhydrite precipitation occurs only in this  
 382 first zone, and a very narrow zone where the montmorillonite remains stable explains the  
 383 modest change in porosity (+1%) over a few centimeters. A third zone where clay minerals  
 384 (illite and montmorillonite) are destabilized but with precipitation of kaolinite, K-feldspar,  
 385 and quartz results in a slight decrease of porosity (1 to 2%). The paragenesis is slightly  
 386 different from *case 1a*, especially concerning the behavior of calcite and anhydrite. In *case 1b*,  
 387 calcite dissolves and anhydrite precipitates in the sulfate-rich caprock water. These  
 388 differences arise from slightly different hypotheses concerning the initial water compositions:  
 389 a complete set of minerals is used in *case 1a* (Table 3) vs. “fast reacting” minerals only in  
 390 *case 1b* (Table 5). In both cases, however, the impact on the porosity profile remains  
 391 comparable.  
 392  
 393



394  
 395 **Figure 4. Porosity and pH profiles in the caprock in the case with buffering capacity of the**  
 396 **reservoir (pH 4.6). Note the difference of zooming in distance.**

397  
 398 The calculations concerning the scenario with the most aggressive water entering the  
 399 caprock show a significant alteration of the caprock at the interface, in comparison to the  
 400 previous case, with a concomitant increase in porosity (Figure 5). The dissolution of  
 401 carbonate (calcite and dolomite) and clay minerals (illite and montmorillonite) is responsible  
 402 for the increase of porosity from the initial 15% to an average value of 85% in the first  
 403 centimeters directly at the interface, even though kaolinite, anhydrite, and quartz precipitate at  
 404 this location. A second front of porosity increase (to almost 60%) reaches 50 centimeters into  
 405 the caprock and corresponds to the front of dissolution of carbonates only (calcite and  
 406 dolomite). The same minerals as in the first front area also precipitate in this zone, except for  
 407 anhydrite which shows a dissolution, up to 15 centimeters, and then a precipitation pattern.  
 408 Porosity and pH variations are closely correlated in these calculations.  
 409



**Figure 5. Porosity and pH profiles in the caprock in the case without buffering capacity of the reservoir (pH 3.4). Note the difference of zooming in distance.**

The results show that in this diffusive case, the caprock alteration can be significant and the impact on porosity greatly depends on the water composition and, in particular, on the pH of the solution. In the case of aggressive water with low pH, all the primary minerals are strongly destabilized directly at the interface between the reservoir and the caprock. A significant increase of porosity is also further predicted but the extent remains limited to 50 centimeters: this sharp porosity front corresponds to the complete dissolution of carbonate minerals (calcite and dolomite) where the other primary minerals are rather preserved, resulting in a porosity increase of about 45%. In the interval between these two fronts, mineral adjustments occur leading to an overall “Z-shaped” porosity profile: dissolution of illite and montmorillonite; precipitation of kaolinite and quartz.

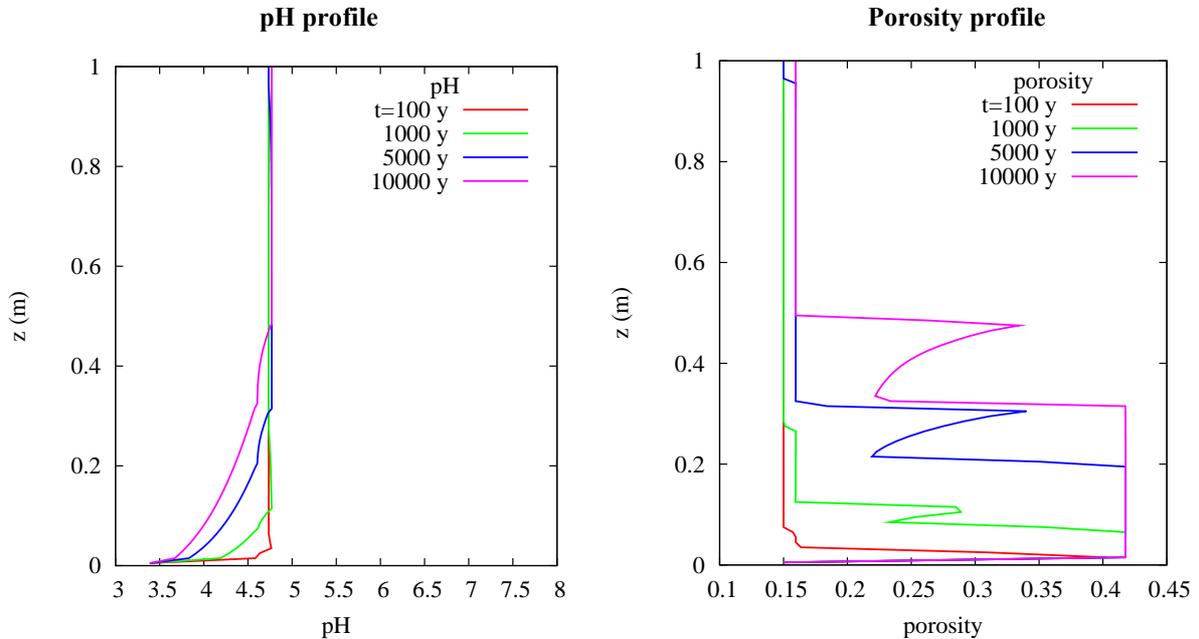
This scenario represents an extreme case where no minerals in the reservoir act as buffer of the pH perturbation, and exemplify the role of the pH of the water at the interface between the reservoir and the caprock. Note also that the mechanical strains that would accommodate at least part of this porosity increase are not taken into account in these calculations. In the other case, the behavior of the system would tend to a slight decrease of the porosity (1%) affecting 3 meters of the caprock. Consequently, if the SC-CO<sub>2</sub> stays confined in the reservoir, the impact of the acidic perturbation on the overall confinement function of a caprock, with a couple of decimeters thickness, will be very low.

***Case 1c: diffusion/slow advection, same initial water composition in reservoir and caprock***

A slightly degraded version of the previous diffusive scenario was simulated. This scenario integrates the effect of a small overpressure in the reservoir compared to the overlying aquifers. This overpressure can result from the injection itself or from the regional hydrological conditions. However, the scenario considers that the overpressure remains low enough so as not to enable a capillary breakthrough of the SC-CO<sub>2</sub>: as a result, only dissolved CO<sub>2</sub> (along with the chemical background of the reservoir pore water) can migrate advectively into the caprock.

An upwards 5 mm/y flow was simulated; this would correspond to a 1 m/m hydraulic head gradient with  $1.6 \cdot 10^{-18} \text{ m}^2$  permeability, a large upper value for the permeability of a deep argillaceous caprock (e.g. Hildenbrand and Kroos, 2003). The diffusive simulations

446 from *case 1b* (low pH) are reproduced, integrating this slow advection. In particular, the sharp  
 447 fronts and the overall "Z-shaped" behavior in the porosity profile are predicted, even though  
 448 they slightly differ in amplitude (maximal increase is 30%). Identical reaction pathways are  
 449 identified, with reaction fronts progressing slightly faster than in the diffusive case. However,  
 450 the progression of the fronts remains limited to the close vicinity of the reservoir/caprock  
 451 interface: Figure 6 shows a migration limited to 0.5 m after 10000 y.  
 452  
 453  
 454



455  
 456  
 457 **Figure 6. Advection/dispersion of CO<sub>2</sub>-rich solution into the caprock: pH and porosity profile as**  
 458 **a function of time.**

459  
 460 The diffusive-advective scenario simulations show that the migration of dissolved CO<sub>2</sub>  
 461 remains confined to the first few decimeters of the caprock after hundreds of years under  
 462 normal hydrological assumptions: homogeneous caprock, quasi natural regional hydraulic  
 463 gradients between the aquifers, no capillary breakthrough.  
 464

465 ***Conclusion on the diffusive case***

466 As a general conclusion to the calculations in homogeneous diffusive/advective  
 467 conditions, we can consider that this case allows us to define a useful reference case for a  
 468 better interpretation of the potential impact of the acidified water diffusion along the caprock.  
 469 Under the assumptions and the initial and boundary conditions considered here, the impact of  
 470 the diffusion of dissolved CO<sub>2</sub> in the caprock is very limited in vertical extension (first  
 471 decimeters to meters after 10,000 yrs). The amplitude depends essentially on the pH of the  
 472 water in the reservoir at the interface with the caprock (a 45% increase in porosity in the first  
 473 50 centimeters in the adverse case of low pH for the reservoir water). In these scenarios, the  
 474 long term consequences of the CO<sub>2</sub> perturbation on the caprock integrity thus appear to be  
 475 small, especially in the context carbonate-dominant storage systems.

476 The same conclusions are reached by other authors in the literature with slightly  
 477 different reactive pathways and consequences on porosity, due to the differences in the initial  
 478 mineralogy. In particular, the presence of plagioclase minerals that tend to dissolve,  
 479 eventually leads to the precipitation of minerals such as dawsonite (Gauss et al., 2005;

480 Gherardi et al., 2007). The intensity of the reactions and porosity changes (up to 10%, with  
481 comparable increasing and decreasing patterns) modeled by these authors are similar to those  
482 observed in this work, even though the temperature was lower ( $\sim 40^{\circ}\text{C}$ ).  
483  
484  
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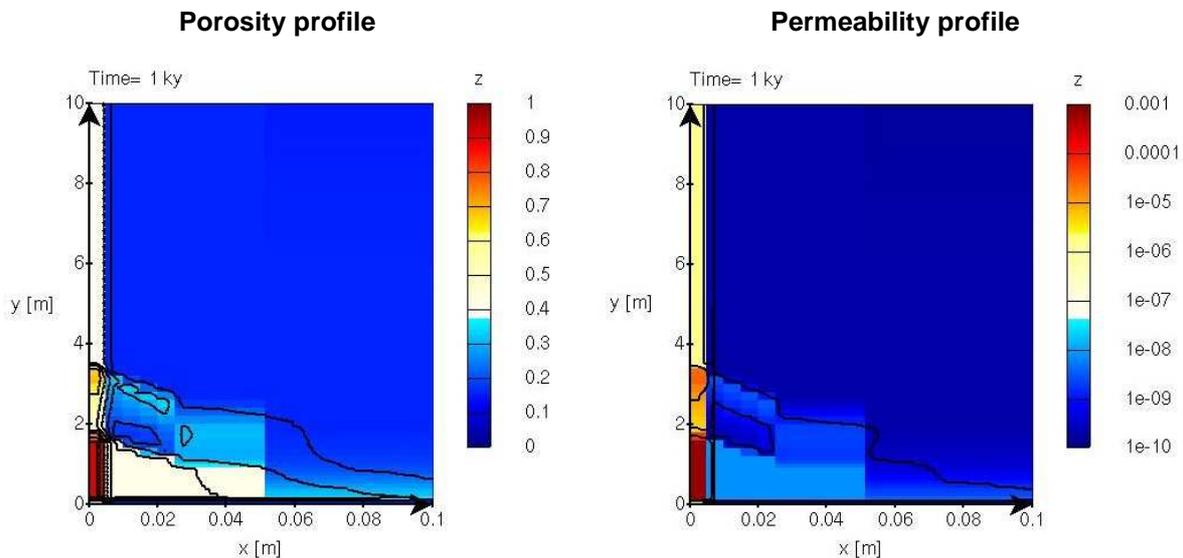
## 486 5.2.2. Caprock with discrete fracture

487

### 488 *Case 2: diffusive/advective transport in discrete fracture*

489 To investigate the effect of local heterogeneities in the caprock, the advective scenario  
490 (*case 1c*) was degraded with the presence of a discrete fracture. The simulation was translated  
491 into vertical 2D, with an explicit fractured zone: a 5 mm wide half-fracture is created into a  
492 homogeneous caprock, on a 10 m high simulation. A 0.1 m/m hydraulic gradient was given,  
493 with a  $1.6 \cdot 10^{-18} \text{ m}^2$  permeability in the bulk matrix, and a permeability  $10^4$  times higher in the  
494 fractured zone (corresponding to 5 mm/y upflow rate). The chemistry in the matrix is taken  
495 identical to that of the previous reference simulations (*case 1c*); the fractured zone is  
496 considered filled up with calcite, with a remaining porosity of 0.4. These conditions constitute  
497 upper values for a realistic system, in agreement with the safety assessment rules of the  
498 exercise.

499 The simulation results (Figure 7) show a degradation front in the bulk matrix far from  
500 the fracture, in agreement with the previous homogeneous simulations in terms of amplitude.  
501 Closer to the fracture, the locally enhanced flow brings more reacting  $\text{CO}_2$  vertically into the  
502 system, so that the reaction front is accelerated (both vertically and horizontally). Inside the  
503 matrix itself, the reaction front is even faster, with a positive feedback of the dissolution of the  
504 calcite, which additionally enhances the fluid flow with the increased permeability in the  
505 reacted area (increased by a factor  $10^3$ ).  
506  
507  
508



509

510 **Figure 7. Distribution of porosity (left) and permeability (right, in m/s) after 1000 years of**  
511 **interactions between the caprock and dissolved  $\text{CO}_2$ . The discrete half-fracture is located at the**  
512 **left side of the domain ( $x < 5 \text{ mm}$ ).**

513

514 The acceleration of the degradation is noticeable in the matrix close to the fracture: 1.5 m  
515 vertically at 1000 y (compared to a 0.4 m/10,000 y in the homogeneous *case 1c*) and in the  
516 fracture itself (4 m after 1,000 years). This is very dependant on the fracture properties:  
517 geometry (aperture), initial permeability, and initial carbonated filling.

518 This is potentially damaging for the sealing properties of the caprock, as opened fractures  
519 could create preferential pathways for SC-CO<sub>2</sub> due to lower local capillary entry pressure. As  
520 a consequence, the reactivity of the carbonated fracture sealing could be furthermore  
521 enhanced, and the reaction front in the fracture even faster.

522 This scenario is different from those already investigated in the literature focusing on the  
523 geochemical effects induced by the advancement of both a CO<sub>2(aq)</sub>-rich aqueous phase and a  
524 free SC-CO<sub>2</sub> gas plume through a highly porous fractured caprock not filled in by secondary  
525 calcite (Gherardi et al., 2007). Under these conditions, the primary mineralogy of the caprock  
526 is predicted to be altered over the entire length of the fracture.

527

528 Generally speaking, the high reactivity of carbonated minerals, which can occur very soon in  
529 the life of the storage (as opposed to kinetically controlled clay mineral reactivity), raises a  
530 potential risk for the caprock integrity, particularly where these mineral phases are dominant  
531 such as in pre-existing fractures. If it appears that CO<sub>2</sub>-saturated water, or SC-CO<sub>2</sub> can  
532 migrate through the caprock, and depending on the transport properties in the fractures, they  
533 could transport the acidic disturbance and potentially open critical pathways for CO<sub>2</sub>.

534

### 535 **5.3. Unsaturated caprock: geochemical interactions with SC-CO<sub>2</sub>**

536

537 These calculations correspond to the scenarios where the SC-CO<sub>2</sub> plume make its way  
538 through the caprock either by overcoming the capillary entry pressure of by migration through  
539 an heterogeneity in the caprock (e.g., by mechanical and/or geochemical reactivation of a  
540 discrete fracture or a network of small cracks).

541

#### 542 **5.3.1. Homogeneous caprock and constant capillary properties**

543

544 **Case 3:** In this section, we will consider the presence of CO<sub>2</sub> as a separate gas phase under  
545 supercritical thermodynamical conditions according to the considered temperature and  
546 pressure of 80°C and 150 bar, respectively. The simulations have been conducted using the  
547 reactive transport code TOUGHREACT. In the following, the term “gas phase” actually  
548 refers to SC-CO<sub>2</sub>.

549 In this work, a shale thickness of 10 m is simulated and we focus on the geochemical impact  
550 of a possible capillary breakthrough of the gas phase in the caprock. Also, later in this work,  
551 we adjust the model of capillary pressure to alter the capillary entry pressure to allow the  
552 breakthrough capillary (see Appendix for a detailed explanation on the role of entry pressure  
553 on capillary trapping). It is important to note that this study does not deal with the estimation  
554 of the capillary entry pressure of the caprock overlying the Dogger in the Paris Basin  
555 geological context and we refer the reader to Chiquet et al. (2007) for a detailed study on this  
556 topic.

557

558

#### 559 *Relative permeability and capillary pressure model*

560 In the framework of the Paris Basin, previous works have been performed to measure  
561 relative permeability and capillary pressure of the so-called Lavoux Limestone considered as  
562 a good analogue for the Dogger reservoir envisaged as a target for geological storage in

563 France (Lombard, 2008, personal communication). In the following, we will assume similar  
 564 trend for relative permeability and capillary pressure model although the cap rock and the  
 565 reservoir pore structure may differ from one geological formation to another. Nevertheless,  
 566 intrinsic permeability and porosity are chosen in agreement with the expected rock texture  
 567 with values of  $10^{-13}\text{m}^2$  and  $10^{-18}\text{m}^2$  assigned to the reservoir and the caprock, respectively.  
 568 André *et al.* (2007) have simulated the measured data using the following models:

$$569 \quad K_{rl} = (S^*)^{0.5} (1 - (1 - (S^*)^{1/m})^m)^2 \quad (1)$$

570 with  $S^* = (S_l - S_{lr}) / (1 - S_{lr})$ ; while the gas relative permeability data have been approximated by  
 571 the following fourth degree polynomial function:

$$572 \quad K_{rg} = 1.3978 - 3.7694S_l + 12.709S_l^2 - 20.642S_l^3 + 10.309S_l^4 \quad (2)$$

573 with  $K_{rl}$  and  $K_{rg}$  the corresponding liquid and gaseous relative permeability phase,  $S_l$  the  
 574 liquid phase saturation,  $S_{lr} = 0.2$  the residual liquid phase saturation and  $m = 0.6$  the van  
 575 Genuchten exponent used in TOUGH2 (see Pruess, 1991). Capillary pressure is approximated  
 576 also with a van Genuchten model described by:

$$577 \quad P_{cap} = -P_o ((S^{**})^{-1/m} - 1)^{1-m} \quad (3)$$

578 with  $P_{cap}$  the capillary pressure,  $S^{**} = (S_l - S_{lr}) / (S_{ls} - S_{lr})$ , and  $P_o$  a pressure coefficient which  
 579 controls the magnitude of the capillary pressure model. In our simulations, we chose  
 580  $S_{ls} = 1.05$ ,  $m = 0.6$  and  $P_o = 100$  kPa to fit with the data of the Lavoux limestone and to allow  
 581 the gas to enter the caprock (see Annexe). Intrinsic permeability values of  $10^{-13}\text{m}^2$  and  $10^{-18}\text{m}^2$   
 582 are assigned to the reservoir and the caprock, respectively.

### 583 *Mesh and boundary conditions*

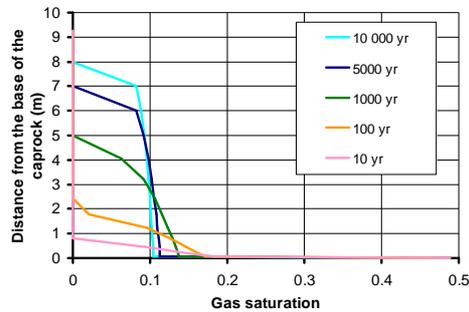
584 The mesh geometry is taken similar to that of the study performed by Xu *et al.* (2005)  
 585 to assess the integrity of a caprock composed with minerals of Texas Gulf Coast sediments.  
 586 The geometry is 1D cartesian along the vertical direction with a  $1 \text{ m}^2$  basal area, and 20 m  
 587 long. The reservoir is represented as a unique cell, 10 m long, while the caprock is meshed  
 588 with 17 cells progressively increasing in thickness from the reservoir-caprock interface to the  
 589 top of the domain (from 0.05 to 1.00 m). The initial state of the system is a hydrostatic profile  
 590 of pressure with uniform temperature at  $80 \text{ C}^\circ$ . A constant pressure boundary was assigned to  
 591 the top limit of the domain.

592 In order to assess the presence of  $\text{CO}_2$  as a separate gas phase and compare this impact  
 593 with the previous simulation, we kept the same initial mineralogical assemblage (Table 3) as  
 594 well as a similar initial brine composition in the caprock (Table 2). On the other hand, the  
 595 reservoir brine is acidified by adding a gas saturation value arbitrarily chosen to be 0.5. The  
 596  $\text{CO}_2$  is assumed to remain in contact with the brine used as a reference (Table 2, column 1)  
 597 composition before injection with an initial pressure of 150 bar.

### 600 *Simulation Results*

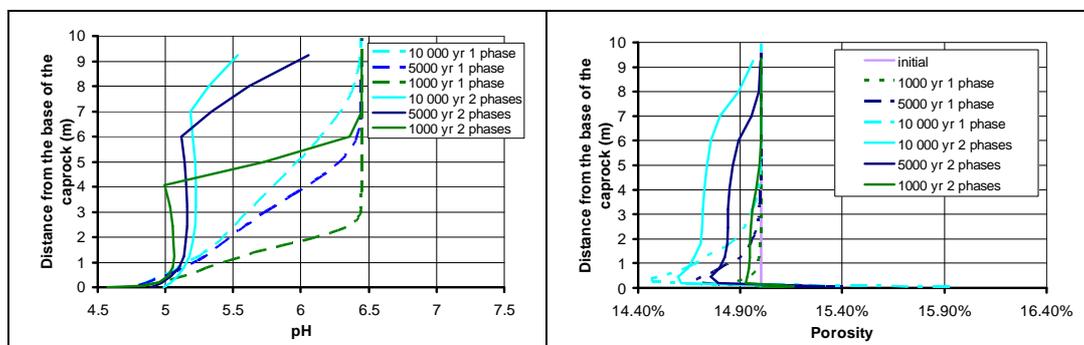
601 Figure 8 shows the simulated upward migration of the gas saturation front through the  
 602 caprock. After ten years the  $\text{CO}_2$  plume has penetrated one meter into the caprock, at 1000  
 603 years it reaches 5 m, and at 10 000 years gas has entered about 8 m into the caprock. This  
 604 upward migration is controlled by both the capillary entry pressure and the intrinsic  
 605 permeability of the caprock. The main process of migration is due to buoyancy effect. Indeed,  
 606 at  $80^\circ\text{C}$  and 150 bar, gas and brine density are equal to  $430$  and  $1020 \text{ kg/m}^3$ , respectively. The

613 pH calculated for the reservoir system equals 4.7 (Figure 9), in good agreement with the  
 614 homologous single phase simulation scenario previously presented (*Case 1a*). The gas front is  
 615 preceded by the dissolution of the gas phase in the brine which acidifies the system. As an  
 616 indication, the results of the single-phase and 2-phase simulations are compared in Figure 9.  
 617



618  
 619 **Figure 8. Gas saturation profiles in the caprock at 10, 100, 1000 and 10 000 years**

620



621  
 622 **Figure 9. pH and porosity profiles at 1000, 5000 and 10 000 years. Comparison with the single**  
 623 **phase case**

624  
 625 After 10 000 years, the simulations predict a small porosity change similar, in  
 626 amplitude, to that observed for the single phase case (*case 1a*) but with a larger extent (Figure  
 627 9). This porosity variation is controlled by the mineral dissolution and precipitation processes  
 628 which are comparable to that of the single phase scenario. Therefore, considering SC-CO<sub>2</sub> in  
 629 the simulation scenario is crucial for the estimate of the capillary trapping in the caprock  
 630 which is controlled by the capillary entry pressure and the thickness of the reservoir. On the  
 631 other hand, if SC-CO<sub>2</sub> penetrates in the caprock, the induced geochemical alteration of the  
 632 shale formation does not differ much from the purely diffusive case prediction, but the  
 633 affected region will occupy a larger extent controlled by the gas plume geometry.

634 Other authors in the literature reached the same conclusions with slightly different  
 635 reactive pathways and consequences on porosity, again due to the presence of plagioclase  
 636 minerals and the precipitation of minerals such as dawsonite (Xu et al., 2005). In contrast, in  
 637 some cases (Johnson et al., 2004; Gherardi et al., 2007), and depending on variable conditions  
 638 of gas saturation and initial mineralogy, major variations in porosity have been predicted in  
 639 the caprock in association with relevant precipitation of other carbonate minerals such as  
 640 magnesite and calcite.

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642 **6. DISCUSSION**

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The results of the modeling show that the injection of CO<sub>2</sub> can potentially have a significant effect on the caprock by changing the mineralogy and changing the porosity due to the dissolution and precipitation of minerals. Although these changes will in turn induced changes in the transport properties of the caprock, the impact is limited to a zone ranging from several decimeters to several meters into the caprock close to the interface with the reservoir depending on whether the SC-CO<sub>2</sub> plume enters into the caprock and/or if fractures are present at this location (Table 6).

	pH		Porosity	
	Minimal value	Impact distance into caprock	Absolute change (- clogging, + opening)	Impact distance into caprock
<b>Saturated homogeneous caprock</b>				
Case 1a (with pH perturbation)	4,8	8 m	- 0.2% + 1.8 %	0.5 - 1.5 m 0.1 - 0.2 m
Case 1a (without pH perturbation)	6,2	7 m	+ 2.2 %	0.1 - 0.2 m
Case 1b (pH 4.6)	4,6	8 m	- 7% - 1%	0.01 m 2.5 m
Case 1b (pH 3.4)	3,4	8 m	+ 70% + 40%	0.05 m 0.5 m
Case 1c (diffusion-advection)	3,4	0.5 m	+ 27% + 18%	0.35 m 0.5 m
<b>Saturated fractured caprock</b>				
Case 2 (1000 y)			+ 25%	1.5 m (vertical) 0.05 m (horizontal)
<b>Unsaturated caprock</b>				
Case 3	4,8	10+ m	- 0.5% + 1 %	0.5 - 6 m 0.1 - 0.2 m

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**Table 6. Overview of minimal pH value and porosity changes with the perturbation distance according to the different scenarios (after 10 000 years or otherwise mentioned in the table).**

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*Heterogeneities in caprock composition and properties*

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The presence of fractures with a different composition and set of transport properties is crucial for the extension of the perturbation. The results obtained with the fracture in the *case 2* show the potential role of other types of heterogeneity, such as the intrinsic heterogeneity of mineralogical composition or the presence of a network of small cracks in the caprock. The geochemical behavior in these systems is intricately coupled with the behavior of the CO<sub>2</sub> plume, through the heterogeneity of capillary properties as demonstrated for instance by Saadatpoor et al. (2009).

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The difference between the reactivity of carbonates and clay minerals is the key to the evolution of the system: short term and short distance for carbonates, long term and long distance for the clay minerals (see *case 1b* and *case 3*). To this regard, one cannot exclude the role of clay transformations such as the illitisation of smectite or illite/smectite interstratified minerals, which is observed in experiments (Crédoz et al., 2010) and potentially represents a further change of properties for the caprock (porosity, permeability, wettability).

671  
672 *Conditions for gas entry into the caprock*

673 In a series of scoping hydraulic/geomechanical calculations (Rohmer and Seyedi, 2009),  
674 the maximal capillary pressure calculated for an industrial-scale injection of CO<sub>2</sub> into a deep  
675 saline aquifer (1 Mt CO<sub>2</sub>/yr) is less than 2 bar at the reservoir-caprock interface, far below the  
676 capillary entry pressure of most caprocks (see results from WP2, Carles et al. 2009; and  
677 Talandier et al. 2006). These results imply that in the conditions of the calculations, the SC-  
678 CO<sub>2</sub> plume would not enter into the caprock in the case of a caprock with homogeneous  
679 properties. This conclusion has to be reassessed in the case where heterogeneities in the  
680 caprock mineral composition could affect the local properties of the rock (such as the  
681 capillary curve or relative permeabilities). Also, for injection periods of 10 to 100 years,  
682 induced plume entry can be triggered by chemical alteration modifying the pore sizes and  
683 structure in the caprock: this scenario was investigated during the project but the model is still  
684 under development and shows numerical instability. The pH changes due to the presence of  
685 CO<sub>2</sub> can also potentially modify the mineral wettability, although this is still a controversial  
686 matter (see e.g. Chiquet et al., 2007; Shah et al. 2008; Fleury et al, *this issue*).

687 During injection, the total overpressure is significant only at the vertical of the injection  
688 point ( $\Delta P = 35$  bar in Rohmer and Seyedi, 2009) after 10 years of injection and drops to a few  
689 bar after the injection stops. These results tend to discard any direct mechanical effect such as  
690 fracturing in the caprock during and after the injection phase. Nevertheless, coupled  
691 mechanical and chemical processes, such as evidenced here, are maximal in the case of long  
692 periods of injection (10 to 100 years) because this would be the typical timescale for  
693 carbonate chemical alteration. This phase may turn out to be critical especially with respect to  
694 the alteration of the sealing properties of the caprock (Bemer and Lombard, 2009).

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696 *Consequences of gas entry into the caprock*

697 The results of the calculations with the complete mineralogical and geochemical system  
698 predict no significant differences between the case with or without gas entry into the caprock  
699 in terms of amplitude of the porosity variations. However, differences are observed on the  
700 extent of the impacted zone within the caprock: several meters in the “gas entry” cases  
701 whereas it is limited to the first decimeters near the reservoir-caprock interface in the cases  
702 without gas entry.

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## 705 **7. CONCLUSION**

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707 The results of the modeling of safety scenarios show that in the normal reference case  
708 (homogeneous rock, diffusion of dissolved CO<sub>2</sub> only) the impact of the reactivity with CO<sub>2</sub>-  
709 rich fluids with the carbonate minerals potentially induces significant changes of porosity.  
710 Nevertheless, this reactivity is limited to the first decimeters of the caprock close to the  
711 interface with the reservoir in 10 000 years and does not lead to any leak from the storage  
712 system.

713 Calculations in “degraded” scenarios show that the migration of the perturbation due to  
714 SC-CO<sub>2</sub> can extend to several meters within the same period of 10 000 years. These scenarios  
715 involve either an alteration of the petrophysical properties of the rock due to the reactivity  
716 with CO<sub>2</sub>-rich solutions (potentially enhancing small existing heterogeneities) or the  
717 reactivation of small cracks or fractures (especially if they are filled with calcite). In both  
718 cases, preferential pathways are created for the migration of CO<sub>2</sub> and positive feedback is  
719 involved, i.e. more migration leads to more reactivity and more alteration of the transport

720 properties. In these scenarios, any convective component occurring as a result of the alteration  
721 will renew and feed the acid perturbation at critical locations (e.g. richer in carbonates  
722 minerals), and can in turn extend the migration distance of CO<sub>2</sub>. One of the key points for the  
723 safety assessment calculations remains the role of these heterogeneities on the behavior of the  
724 storage system. We observe in this work the great influence of structural heterogeneities. The  
725 quantitative assessment of such an impact on the global safety remains an important challenge.  
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727 To complete a safety assessment for a specific site, a set of conservative (penalizing)  
728 parameters should be adopted for adverse scenarios (such as in *case 1b*) to evaluate the  
729 quantity of CO<sub>2</sub> release from the storage system and the probability of occurrence of these  
730 scenarios should be evaluated from data available from this site (Bildstein *et al.* 2009). The  
731 challenge for the future will be to structure and apply the safety assessment methodology with  
732 an operational finality, in order to support the transition step to carbon geological storage  
733 projects at the industrial scale.  
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## 735 8. ACKNOWLEDGEMENTS

736 The authors acknowledge funding from the ANR (the French National Science  
737 Foundation) in the framework of the “Géocarbonate” projects (Injectivité, Carbonation,  
738 PICOREF, CRISCO2) and fruitful discussion with the other partners in these projects. This  
739 project also benefited greatly from the scientific exchanges taking place in the working group  
740 “Modeling at the site scale” animated by E. Brosse (IFP).  
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## 742 9. APPENDIX: ROLE OF ENTRY PRESSURE ON CAPILLARY TRAPPING

743 Figure 10 shows the pressure fields of water and carbon dioxide in a geological  
744 storage in aquifer located at a depth  $H$  underneath the surface. Once injected, CO<sub>2</sub> occupies  
745 the reservoir pore space available initially filled with water. Due to gravity effects, the storage  
746 area, containing both CO<sub>2</sub> and residual water, is located at the top of the aquifer and has a  
747 thickness  $h$ . Point **A** is located at the interface between reservoir and caprock and is part of the  
748 caprock domain. Point **A'** is also located at the interface between reservoir and caprock but  
749 belongs to the reservoir domain. Point **B** is located at the gas-water contact. At this interface  
750 the capillary equilibrium is reached and the capillary pressure is zero. In other words in **B**:

$$751 \quad P_{B,CO_2} = P_{B,H_2O}$$

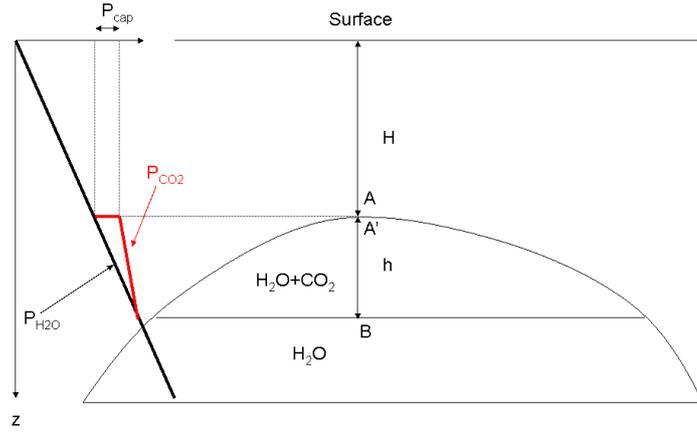
752 In **A**, the water pressure is equal to the water column in the caprock:

$$753 \quad P_{A,H_2O} = \rho_{H_2O} g H$$

754 In **A'**, the CO<sub>2</sub> pressure  $P_{A',CO_2}$  is equal to:

$$755 \quad P_{A',CO_2} = P_{A,H_2O} + P_{cap}$$

756 with  $P_{cap}$ , the capillary pressure defined as the difference between gas pressure and water  
757 pressure.  
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**Figure 10 : CO<sub>2</sub> and H<sub>2</sub>O pressure evolution with depth in aquifer storage**

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In the context of geological storage, we define the capillary entry pressure as a capillary pressure threshold representing the resistance of porous network saturated with water to the penetration of carbon dioxide. The capillary entry pressure  $P_{ce}$  is intrinsic to the nature of the pore network constituting the clay caprock. According to the Laplace law, it can be expressed as:

$$P_{ce} \sim 2\sigma \cos(\phi) / R,$$

with  $\sigma$  the inter-facial tension between the liquid and gas,  $\phi$  the angle reflecting the ability of the liquid to spread over a surface by wettability, and  $R$  the radius of the largest pores (Chiquet et al 2007).

The capillary entry pressure is the value of capillary pressure for a zero gas saturation (Figure 11). The van Genuchten model used in the code Toughreact adjusts the capillary entry pressure using the liquid parameter saturation  $S_{ls}$ . For  $S_{ls} > 1$ , the model of capillary pressure becomes non zero for  $S_g = 0$  (Figure 11).

If the capillary pressure at the caprock-reservoir interface exceeds the capillary entry pressure ( $P_{cap-interface} > P_{ce}$ ), then the CO<sub>2</sub> can penetrate the caprock.

The capillary pressure in the reservoir-caprock interface  $P_{cap-interface}$  is expressed as:

$$P_{cap-interface} = P_{A',CO_2} - P_{A,H_2O}$$

Assuming no change with depth in both CO<sub>2</sub> and H<sub>2</sub>O densities, we can write:

$$P_{A',CO_2} = P_B - \rho_{CO_2}gh$$

$$P_{A,H_2O} = \rho_{H_2O}gH$$

On the other hand, considering that:

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$$P_B = \rho_{H_2O} g (H+h)$$

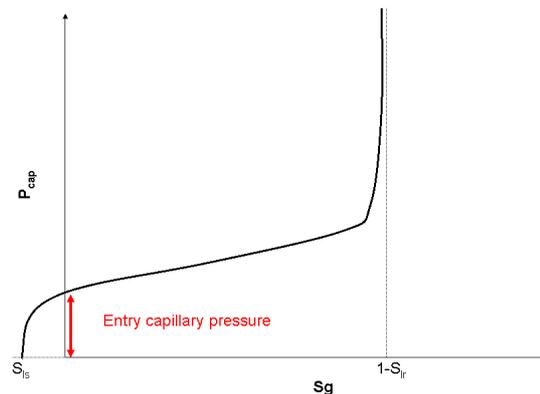
We finally obtain:

$$P_{cap-interface} = (\rho_{H_2O} - \rho_{CO_2}) g h$$

This corresponds to the buoyancy effect induced by density difference between liquid and gas.

Thus, for the gas phase to penetrate the caprock, we must verify that:

$$P_{ce} < (\rho_{CO_2} - \rho_{H_2O}) g h$$



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**Figure 11 : Representation of the entry capillary pressure using the van Genuchten model as described in the code TOUGHREACT**

In our system, we assume a 10 m thickness for the reservoir. At 150 bar and 80°C, the density of CO<sub>2</sub> is about 430 kg/m<sup>3</sup>, while the density of water is 1018 kg/m<sup>3</sup>, which leads for a 10 m thick reservoir to a value of  $P_{cap-interface}$  of 57 kPa (0.57 bar). Using a value of  $P_o = 100$  kPa in Eq(3) yields a  $P_{ce}$  of 54 kPa slightly below  $P_{cap-interface}$  and then allowing CO<sub>2</sub> to penetrate the caprock.

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